



Scientific Committee on Consumer Safety

SCCS

Opinion on Nitrosamines and Secondary Amines
in Cosmetic Products

The SCCS adopted this opinion at its 14th plenary
of 27 March 2012

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Three independent non-food Scientific Committees provide the Commission with the scientific advice it needs when preparing policy and proposals relating to consumer safety, public health and the environment. The Committees also draw the Commission's attention to the new or emerging problems which may pose an actual or potential threat.

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In addition, the Commission relies upon the work of the European Food Safety Authority (EFSA), the European Medicines Agency (EMA), the European Centre for Disease prevention and Control (ECDC) and the European Chemicals Agency (ECHA).

SCCS

The Committee shall provide opinions on questions concerning all types of health and safety risks (notably chemical, biological, mechanical and other physical risks) of non-food consumer products (for example: cosmetic products and their ingredients, toys, textiles, clothing, personal care and household products such as detergents, etc.) and services (for example: tattooing, artificial sun tanning, etc.).

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http://ec.europa.eu/health/scientific_committees/environmental_risks/index_en.htm

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All Declarations of working group members are available at the following webpage:
http://ec.europa.eu/health/scientific_committees/environmental_risks/members_wg/index_en.htm

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1. BACKGROUND

The possible health risks that may be associated with the presence of nitrosamines in cosmetic products are regulated under the Cosmetics Directive (76/768/EEC) under the following entries of Annexes II (total ban) and III (restricted use):

Annex II:

410. Nitrosamines

411. Secondary alkyl- and alkanolamines and their salts

Annex III:

60. Fatty acid dialkylamides and dialkanolamides. Maximum secondary amine content in the finished product: 0.5%. Maximum secondary amine content in the raw materials: 5% (applies). Maximum nitrosamine content 50 µg/kg. Should not be used with nitrosating systems and should be kept in nitrite-free environments.

61. Monoalkylamines, monoalkanolamines and their salts. Maximum secondary amine content in the finished product: 0.5%. Minimum raw material purity: 99%. Maximum secondary amine content in the raw materials: 0.5% (applies). Maximum nitrosamine content 50 µg/kg. Should not be used with nitrosating systems and should be kept in nitrite-free environments.

62. Trialkylamines, trialkanolamines and their salts. Maximum level in non-rinse-off products: 2.5%; Minimum raw material purity: 99%. Maximum secondary amine content in the raw materials: 0.5% (applies). Maximum nitrosamine content 50 µg/kg. Should not be used with nitrosating systems and should be kept in nitrite-free environments.

(a) 2.5% (a) (b)

The scientific basis for the above provisions are the relevant opinions of the Scientific Committee on Cosmetology (SCC/023/92) and the Scientific Committee on Cosmetics and Non Food Consumer Products (SCCNFP/0110/99 and SCCNFP/0462/01).

The implementation of the above provisions in the market place has been difficult as the chemical terms used in the above entries are relatively generic and allow for different interpretations among economic operators and public authorities, ranging from those which consider that all chemical structures which contain a secondary amine group are covered by the above entries to those who take a more practical approach to consider only those secondary amines with unsubstituted or monosubstituted alkyl and/or alkanoyl groups.

According to the Cosmetics industry entry 411 of Annex II is to be interpreted in the following manner:

- A secondary alkylamine is one in which precisely two alkyl groups are attached to a nitrogen atom, with no further substitution of either the nitrogen or the alkyl groups. However, even this interpretation is ambiguous since a secondary alkylamine could be an amine carrying one or more secondary alkyl groups.
- A secondary alkanolamine is one in which the nitrogen atom carries precisely two monohydroxyalkyl groups, with no further substitution.

On the other hand, the Norwegian Competent Authority has asked the Commission for further assistance in the evaluation of the interpretation of entry 411 of Annex II concerning some polyamines like spermidine (CAS 334-50-9), Gerotine (CAS 71-44-3) and dipropylentriamine (CAS 56-18-8) which are used as ingredients in cosmetic products

marketed by a Norwegian company. These are secondary aminoalkylamines meaning that strictly speaking they are neither alkylamines, nor are they secondary alkanolamines and hence should not be covered by entry 411 of Annex II although by definition (as secondary amines) can give rise to nitrosamines.

Dipropylenetriamine is a primary amine which contains a secondary group but chemically speaking it is neither a primary alkylamine nor a primary alkanolamine. Hence in the strict sense it is outside the scope of Annex III, part 1, entry 61, yet from the safety point of view this is rather illogical since it is a chemical which under appropriate conditions (nitrosating agents, heat etc) will most certainly give rise to its nitrosamine upon nitrosation of its secondary amine group. Other polyamines which potentially create similar interpretation challenges like Azamethonium (CAS 60-30-0), Pentamethonium (CAS 2365-25-5), and Hexamethonium (CAS 55-97-0) are banned under entries 121, 120 and 124 of Annex II for different reasons.

The situation is similar for the fatty acid dialkylamides and dialkanolamides. Here it is the fatty acids moiety that needs to be considered. They are normally understood to mean long-chain alkanolic or alkenolic acids, and would only include limited branching; stearic acid, isostearic acid, and oleic acid would all be considered fatty acids. When the alkyl chain is further substituted or if it is not linear, they do not 'fit' the standard definition of a fatty acid as understood in practice although chemically they are fatty acids. Thus cocoyl sarcosine is not considered to be fatty acid, and cocoyl sarcosinamide DEA is not considered to be a fatty acid dialkanolamide, and therefore not subject to the restrictions by Annex III, part 1, 60, yet they can both give rise to nitrosamines. Of course, even though these substances are not subject to any specific restrictions, they are still bound by annex II/410 and 411, and any secondary alkylamine or alkanolamine used in the manufacture must be absent from the finished product except as technically unavoidable levels, and as long as the finished product complies with Article 2 of the Cosmetics Directive.

The above examples clearly demonstrate the need for a proper definition of the relevant entries in Annexes II and III taking into account the scientific basis of the previous SCC and SCCNFP opinions and the evolving scientific knowledge. In those opinions and as documented in the scientific literature available in the public domain, the scientific basis for the entries into the Annexes II and III of the Directive was the relative genotoxic carcinogenicity potential of certain well studied nitrosamines which are (or were at the time) most commonly found as contaminants in cosmetic and consumer products containing mono-, di- and trialkyl amines. At the same time, the SCC and SCCNFP opinions and the evolving scientific knowledge make it absolutely clear that no blanket statements and generalisations should be used to manage the potential risks of nitrosamines because 1) there is to date no evidence that all nitrosamines that can be formed when secondary amine group(s) present in a given chemical have carcinogenic potential. Besides the fact that for most of them there is simply no data, other factors may play a role in this such as the chemical stability of the nitrosamine, the special arrangements of the molecules which may hinder its genotoxic potential, etc; and 2) not all nitrosamines which have been studied to date exhibit the same inherent genotoxic and carcinogenic potential.

Finally, a definition of the wording used in annex III entries 60-62 of "nitrosating systems" is needed as this has also led to different interpretations.

2. TERMS OF REFERENCE

Taking the above together and in light of the previous SCC, SCCNFP opinions on the subject and the currently available data on the genotoxicity/carcinogenicity of nitrosamines, the SCCS is asked to:

Elaborate an opinion on the potential risks to human health by the presence in cosmetics of nitrosamines or of chemicals with secondary amine groups which may give rise to N-nitroso compounds, to provide guidance to the Commission in revising the relevant entries of Annexes II and III of the Cosmetics directive (76/768/EEC). To this end, the SCCS should:

- 1) Identify chemical classes that can give rise to nitrosamines.
- 2) Provide a definition (or provide a generic definition) of the substances regulated in Annex II 411 and Annex III 60-62, i.e. secondary alkylamine and secondary alkanolamine, fatty acid dialkylamides and dialkanolamides and mono- and tri-alkylamines and alkanolamines.
- 3) Comment on the possibility to group chemicals and/or chemical classes with respect to their reactivity towards nitrosating agents and their propensity to give rise to nitrosamines. Identify chemicals or groups/classes for which such grouping with respect to nitrosation may not be possible and case-by-case assessments need to be made.
- 4) Identify the factors/conditions that may influence/enhance /inhibit the formation of nitrosamines i.e. N-nitroso compounds (e.g. N-Nitroso-oxazolidines), such as nitrogen oxides, nitrite, preservatives, catalysts (e.g. formaldehyde) or others. Provide a clear definition for nitrosating systems. Clarification is required to address whether a nitrosating agent or a nitrosating system should be basis for the regulation of nitrosamine formation in cosmetic ingredients and cosmetic formulations.
- 5) List the nitrosamines found in cosmetics and advise the Commission of approaches to rank nitrosamines that may occur in cosmetics with respect to their carcinogenic potency.
- 6) Is there a way to identify chemical classes, and ranking them in terms of their propensity to give rise to carcinogenic nitrosamines and their potency? Inversely, is there a way to relate the carcinogenic potential of nitrosamines formed with the parent chemical class?
- 7) Comment on the levels of 50 µg nitrosamine/ kg *as set out currently in the Annexes of Directive 76/768/EEC*. Should it apply to finished products or to raw materials? Should it be considered for all nitrosamines potentially formed? Should it be modified, following the ranking of carcinogenic potency of nitrosamines in question? Comment on the "maximum secondary amine content (5% in raw materials and 0.5% in finished products)".
- 8) On the basis of the answers above SCCS to pronounce itself
 - on the specific cases of spermidine (CAS 334-50-9), gerotine (CAS 71-44-3) and dipropylentriamine (CAS 56-18-8);
 - on the "Maximum secondary amine content: 5% (applies to raw materials)" and that "Maximum secondary amine content: 0.5%" in the finished cosmetic products" for the Fatty acid dialkylamides and dialkanolamines listed in entry 60 of Annex III, part I.

PUBLIC CONSULTATION

A public consultation of the preliminary report of this opinion took place from 19 December 2011 to 10 February 2012. Three contributions were received during the consultation period.

In evaluating the responses from the consultation, submitted material was considered under the following conditions:

1. Concerned issues were under the remit of the Scientific Committees
2. Referred directly to the content of the report and was related to the issues addressed in the report,
3. Contained specific comments and suggestions on the scientific basis of the opinion,
4. Has the potential to add to the preliminary opinion of Scientific Committee.

Each contribution which met these criteria was carefully considered by the Working Group and when appropriate the text of the opinion was modified.

3. INTRODUCTION

N-Nitroso compounds (NOC) are amongst the most potent carcinogens. More than 300 of these compounds have been tested in about 40 mammalian animal species, including subhuman primates. NOC have been applied by oral, subcutaneous, intraperitoneal, intravenous or subcutaneous injection, as well as by local instillation into the trachea, the urinary bladder or the colon and rectum. They were also applied topically to the skin of experimental animals. Irrespective of the mode of application, no species has been found to be resistant against the carcinogenic efficacy of these compounds. It has been also shown that their metabolism in animals is similar to the metabolism in human tissues. The probability therefore is high that N-Nitroso compounds (NOCs) may be carcinogenic in humans as well. Moreover, the likelihood of a new NOC of as yet unknown biological activity to be genotoxic/ mutagenic/ carcinogenic is very high. NOCs comprise nitrosamines and nitrosamides. Whereas nitrosamines are to be metabolically activated, mostly by cytochrome P450 dependent enzymes, nitrosamides do not require metabolic activation for genotoxic/ carcinogenic activity. In cosmetics only few NOCs have been identified. Mainly two nitrosamines, N-nitrosodiethanolamine (NDELA) and N-nitrosobis (2-hydroxypropyl)amine (NBHPA) have been found.

4. OPINION

4.1. Chemical classes that can give rise to nitrosamines

Compounds / cosmetic constituents considered precursors of relevance for generation of N-Nitroso compounds, i.e. nitrosamines in raw materials and finished cosmetics.

Primary, secondary and tertiary amines can all be nitrosated to generate nitrosamines. The secondary amines in general are the most reactive compounds towards nitrosating agents, generating nitrosamines.

Primary alkyl amines react with nitrosating agents to give short-lived, highly reactive diazonium ions. These reactive intermediates decompose to give molecular nitrogen by substitution, elimination, and molecular rearrangement pathways. However, nitrosamines occasionally arise from secondary processes (see section 4.3.1. below).

Any compound containing the secondary amine functional group is expected to react with nitrosating agents to produce a NOC. The extent of nitrosamine formation will depend upon

the structural features of the amine that influence the rates of these transformations and the concentration of the nitrosating agent.

Nitrosamines may also be formed from tertiary amines (Williams, 2004; Mirvish, 1975b; Loeppky et al, 1990). The transformation requires the cleavage of the carbon-nitrogen bond of one of the alkyl groups attached to the nitrogen atom. The rates and the nature of these processes depend significantly on the structure of the tertiary amine (Loeppky et al, 1990). Some tertiary amines have unusually high reactivity towards nitrosation, related to the presence of special structural features (Loeppky et al, 1990).

Amides and related carbonyl derivatives of amines and of ammonia can react with nitrosating agents to generate NOCs (Williams, 2004; Mirvish, 1975b). Because of the electron withdrawing properties of the carbonyl group attached to the nitrogen, the reactivity of amides toward common nitrosating agents is low. Primary amides react with nitrous acid to give the parent acid of the amide and N₂, and can be used as nitrous acid traps. Amides of dialkanolamines can give rise to the corresponding nitrosamines by reaction with ionic nitrites (Loeppky et al, 1982, 1984).

4.2. Definitions of substances regulated

Amines are classified as primary, secondary or tertiary, corresponding to one, two, or three alkyl or aryl groups bonded to nitrogen (Wade, L.G., Jr. Organic Chemistry, 6th edition 2006, Pearson Prentice Hall, ISBN 0-13-147871-0). A **secondary** amine may also be regarded as a derivative of ammonia in which two of the hydrogen atoms have been replaced by two organic groups (which may be equal or different from each other and may carry further substituents), linked to the nitrogen by single bonds to the carbon atoms. These carbon atoms cannot be doubly or triply bonded, or bound to other hetero atoms, such as oxygen or nitrogen.

A **secondary** alkylamine is a derivative of ammonia in which two of the hydrogen atoms have been replaced by two alkyl groups linked to the nitrogen by single bonds to the carbon atoms.

A **secondary alkanolamine** is a derivative of ammonia where two of the hydrogen atoms are replaced by alkanol groups. The hydroxyl groups should not be substituted.

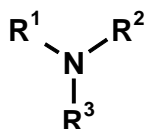
A **fatty acid dialkylamide** is a derivative of a fatty acid amide, where the alkyl fatty acid moieties could be a branched chain or unsaturated.

A **fatty acid dialkanolamide** is a fatty acid derivative of a dialkanolamide, where the alkyl fatty acid moiety could be a branched chain or unsaturated.

A **monoalkylamine** (primary amine) is a derivative of ammonia, where one hydrogen atom is replaced by an alkyl group. A **monoalkanolamine** is a derivative of a monoalkylamine, where the alkyl group is replaced by an alkanol group.

A **trialkylamine** (tertiary amine) is a derivative of ammonia, where hydrogen atoms are replaced by alkyl groups. A **trialkanolamine** is a derivative of trialkylamine, where three of the alkyl groups are replaced by an alkanol group.

Structures



Amines

R^1, R^2, R^3	= H	ammonia
R^1	= organic group	primary amine
R^1, R^2	= organic group	secondary amine
R^1, R^2, R^3	= organic group	tertiary amine

Organic group: alkyl, alkanol, alkylamine, aryl, aryl-alkyl

Amides: R^1 = organic acid residue

R^2, R^3	= H	primary amide
R^2	= organic group	} secondary amide
R^3	= H	
R^2, R^3	= organic group	tertiary amide

4.3. Grouping of chemicals with respect their nitrosating reactivity

4.3.1. Grouping of chemicals

Comment on the possibility to group chemicals and/or chemical classes with respect to their reactivity towards nitrosating agents and their propensity to give rise to nitrosamines.

General considerations:

Nitrosamines may inadvertently be formed during the process of manufacture of commercial preparations, often during the formulation process through the action of unsuspected added ingredients or environmental conditions. While in the case of cosmetics and personal care items, these transformations rarely involve acidic nitrosation, the classical method for preparation of nitrosamines; it is instructive to review some aspects of the acidic nitrosation process in order to reveal how substrate structure modifies reactivity.

There have been numerous kinetic investigations of secondary amine acidic nitrosation, and the subject has been well reviewed from several perspectives (Williams, 2004; Mirvish, 1975b). The following slightly edited quotation provides a succinct summary (Eisenbrand, 1990).

"The most important factors determining the kinetics of amine nitrosation are represented by the acid-base equilibria of the amine and of nitrous acid. The nitrosating agent N_2O_3 reacts with the free electron pair of the unprotonated amine to the corresponding N-nitroso compound, displacing NO_2 .

Since the nitrous acid/nitrite equilibrium has a pKa of 3.4, at a pH of 3 about half of the nitrite will be in the reactive, protonated form. Furthermore, since two molecules of HNO_2 form the nitrosating intermediate, N_2O_3 , the rate of nitrosation follows second order kinetics with respect to nitrous acid (Mirvish, 1975b). Low pH favours N_2O_3 formation, but reduces concentration of the unprotonated amine, the reactive species available for nitrosation. The reverse is true for higher pH values. For most alkylamines therefore the pH-dependence of the reaction rate represents a bell-shaped curve, with optimal rates at pH 3-4 and mostly steep slopes above and below this optimal value (Mirvish, 1970)".

Moreover, it has been established that the intrinsic rate constants for the reaction of the free amine with N_2O_3 for a number of amines are remarkably similar (Williams, 2004; Mirvish, 1975b). Consider, for example, the nitrosation rate constants ($\times 10^8 M^{-1}sec^{-1}$) for morpholine and piperidine are 2.2 and 1.8, respectively. Yet, their pKa values (8.7 and 11.2, respectively) and experimentally observed nitrosation rates are vastly different. Morpholine, the less basic amine, reacts at least 316 times faster at pH 3.4 under

comparable conditions. This comparison and numerous others point to the fact that basicity is the major factor that controls the nitrosation rate of amines.

Simple methods for estimating nitrosation rates based on the pKa of the ammonium ion of the secondary amine have been developed (Eisenbrand, 1990; Mirvish, 1975b). The basicity of an amine is determined by its structure. Piperidine has a six-membered ring structure containing a single N-H in the ring. Morpholine has a similar structure, except that the CH₂ directly opposite to nitrogen, has been replaced by an oxygen atom. This structural change diminishes the basicity by 2.5 pKa units. The change results from the electron withdrawing capacity of the oxygen atom in comparison to the less electronegative carbon of piperidine. In general, compared to a normal alkyl chain, substitution of CH₂ by a more electronegative oxygen or, to a lesser extent, a nitrogen atom, will decrease the basicity and increase the nitrosation rate of the amine. Since the pKa values of many amines are available, nitrosation rates can be relatively easily predicted under specified conditions.

Replacement of one of the alkyl groups of a secondary amine by a benzene ring reduces the amines basicity by approximately 5 pKa units. Substituents in the benzene ring can have a powerful effect on the nitrosation rate of the secondary amine, because they affect both the amine basicity and the intrinsic rate constant for nitrosation (Williams, 2004). So here, the simple relationship between basicity and nitrosation rate breaks down. Electron withdrawing substituents in the benzene ring decrease basicity, but they also decrease the intrinsic rate constant for nitrosation. Electron donating substituents in the benzene ring have the opposite effect and overall increase nitrosation rates.

Although not common, steric crowding adjacent to the nitrogen atom of the secondary amine reduces nitrosation rates. This effect was demonstrated for piperidines methylated at the 2- and 6-positions (Jones et al, 1974). Stereo electronic effects are quite important in determining tertiary amine nitrosation rate, see Section below

Groupings

Primary amines

Primary alkyl amines react with nitrosating agents to give short-lived, highly reactive diazonium ions (the species also formed by metabolic activation of nitrosamines). These reactive intermediates decompose to give molecular nitrogen by substitution, elimination, and molecular rearrangement pathways. For example, the nitrosation of ethanolamine in acetic acid produces ethylene glycol, 2-hydroxyethyl acetate, and acetaldehyde (Park et al, 1998). The generated acetaldehyde can react with ethanolamine to form 5-methyloxazolidine which may be nitrosated to the corresponding NOC. Depending on the primary alkanolamine, other NOC may occasionally arise from such secondary processes (see section 4.4.5 below).

Spermidine, a polyamine containing both primary and secondary amino groups, reacts with nitrosating agents to give N-nitrosopyrrolidine and other nitrosamines. It is likely formed from the reaction of a diazonium ion derived from one of primary amino groups with the secondary amine to close the ring (see section 4.8. (Bills et al, 1973; Huxel et al, 1974).

Secondary amines, alkylamines and alkanolamines:

Any compound containing the secondary amine functional group is expected to react with nitrosating agents to produce a nitrosamine. The extent of nitrosamine formation will depend upon the concentrations of the amine and the nitrosating agent as well as structural features of the amine influencing reaction rates as discussed above.

Tertiary amines

Nitrosamines may also be formed from tertiary amines. The transformation requires the cleavage of the carbon-nitrogen bond of one of the alkyl groups attached to the nitrogen atom. The rates and the nature of these processes depend significantly on the structure of the tertiary amine (Williams, 2004; Mirvish, 1975b; Loepky et al, 1990). In the case of tri-alkyl amines, the scope, mechanism (Smith et al, 1967), and kinetics have been well

studied (Mirvish, 1975b; Smith et al, 1967; Gowenlock et al, 1979). A number of nitrosating agents are capable of affecting this transformation, but the reaction is slow compared to secondary amine nitrosation and often requires a significant molar excess of a nitrosating agent and temperatures above 60°C to occur at a reasonable rate. Because of these facts, tertiary alkanolamines and other tri-alkyl amines may be used in many commercial formulations of interest without concern for nitrosamine formation.

Aromatic di-alkyl amines are much more reactive toward nitrosation than are the tri-alkyl amines (Williams, 2004; Loeppky et al, 1983, 1991, 1998). The products of these reactions are arylalkyl nitrosamines. Only those compounds containing a powerful electron withdrawing group in the benzene ring do not react at an appreciable rate. Sun screen preparations containing esters of 4-dimethylaminobenzoic acid have been shown to contain N-nitroso compounds produced during formulation (Loeppky et al, 1991). Significant advances have been made in the understanding of the chemical details of these transformations (Williams, 2004; Loeppky et al, 1983, 1991, 1998; Teuten et al, 2005).

Some tertiary amines have unusually high reactivity towards nitrosation because of the presence of special structural features (Loeppky et al, 1990). Some of these have been reviewed (Williams, 2004; Mirvish, 1975b; Loeppky et al, 1990). However, because the chemistry is hard to generalize, it is recommended that manufacturers of products containing aromatic di-alkyl amines use the Apparent Total N-nitroso group Content (ATNC) procedure defined and discussed in the COLIPA technical document (COLIPA, 2009) to demonstrate that their products are nitrosamine free.

Quaternary ammonium compounds:

Also quaternary ammonium compounds may be of relevance with respect to NOC formation, as may be deduced from reports on acute liver toxicity of combinations of nitrite and certain quaternary ammonium compounds (Maduagwu, 1985). Moreover, amine nitrosation by nitrite has been reported to be catalyzed in presence of the cationic surfactant decyltrimethylammonium bromide and other micelle forming agents (Okun et al, 1977), suggesting amine impurities in such surfactants being prone to enhanced nitrosation risk.

Amides. The reactivity of amides and related carbonyl derivatives of amines and of ammonia toward nitrosating agents has been reviewed recently (Williams, 2004), but already a 1975 review (Mirvish, 1975b) contains much pertinent information. Because of the electron withdrawing properties of the carbonyl group attached to the nitrogen, the reactivity of amides toward common nitrosating agents is low. Primary amides react with nitrous acid to give the parent acid of the amide and N₂, and can be used as nitrous acid traps. Powerful nitrosating agents such as fuming nitric acid, N₂O₄, and NOBF₄ are required to convert secondary amides into their N-nitroso derivatives. While these same reagents have little effect on tertiary amides, ionic nitrite can convert them into nitrosamines. Amides of dialkanolamines are especially subject to nitrosation by ionic nitrites (Loeppky et al, 1982, 1984) See also section 4.4.3.

N-alkyl-carbamates are more readily nitrosated than amides (Mirvish, 1975b). In N-alkylureas, the presence of the second carbamoyl nitrogen brings about greatly enhanced nucleophilicity and such compounds are very easily nitrosated, forming nitrosoureas, most of which are strong carcinogens.

4.3.2. Cases where grouping with respect to their reactivity to nitrosating agents is not possible

Identify chemicals or groups/classes for which such grouping may not be possible and case-by-case assessments need to be made.

This group may include dyes and other nitrogen-containing additives that are difficult to classify. It is recommended that manufacturers of products containing these substances use the Apparent Total N-nitroso group Content (ATNC) procedure defined and discussed in the COLIPA technical document (COLIPA, 2009) to demonstrate that their products are nitrosamine free. When applying the ATNC methodology care should be taken to verify its

reliability and to exclude erroneous results, which may be produced under specific circumstances, for instance in the case of some hair dyes. Since it is based on liberation and detection of NO, response it is not confined to N-Nitroso groups, but also, for example to S-nitroso as well as C- and N-nitro groups.

Specific case of hair dyes ingredients:

Primary, secondary and tertiary amines are ingredients of both oxidative and non-oxidative hair dye formulations. These are used in concentrations of 0.1-3% (on head). In general, these amines invariably are attached to an aromatic, heteroaromatic or polycyclic ring system that may carry hydroxy-, amino-, alkylamino-, alkoxy- and other substituents.

Monosubstituted nitrogens attached to aromatic rings are very easily nitrosated under nitrosating conditions. Examples for compounds carrying easily nitrosatable nitrogens are, for instance, COLIPA numbers A9, 22, 25, 31, 84, 98, 130, 138, B5, 70, 34, 36, 37 and similar compounds (table 1).

Disubstituted nitrogens attached to aromatic rings may generate NOC as well. In general the rate of NOC formation is expected to be lower for tertiary amines, as compared to secondary amines, but there are exceptions. When the N-nitrosatable group is attached to an electron-rich aromatic ring, reactivity towards nitrosating agents to give nitrosamines by dealkylation is greatly increased. Furthermore, certain compounds may also react through a radical cation mechanism, as has been shown for compounds with similarity to A50 (Williams, 2004; Eisenbrand, 1990). A50 is expected to produce a nitrosamine and a nitro compound in a radical cation mechanism, encompassing competition of N-dealkylation with ring nitration and with diazotization of the primary amino group. There are other tertiary amines on the list that may be conceptualized to react similarly, for example: A121, B31, B37, B73, B77 (see table 1 below)

It is difficult to predict the precise product ratios and rates. Of note, in presence of an aromatic nitro substituent in a hair dye constituent, it is important to ensure that no nitrosating potential is exerted.

C117 is a compound that in addition to carrying a nitrosatable amino group as part of an aminoanthraquinone ring system also carries an N-Methyl-morpholinium substituent that under nitrosating conditions may give rise to N-Nitrosomorpholine.

B70, represents a urea structure, monosubstituted with a 2-(4 Nitrophenylamino) ethyl substituent. It has two easily nitrosatable nitrogens, one of which expected to form a directly acting NOC. Since during the oxidative hair dyeing process mostly basic conditions prevail, this NOC is expected to rapidly decompose.

A154 represents an N-substituted diaminopyrazole. In the presence of nitrosating agents it is expected to form a diazonium intermediate that may generate the corresponding 1,2,3-triazole by ring closure reaction.

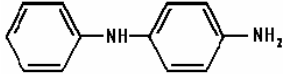
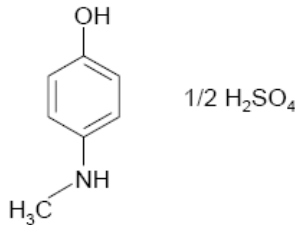
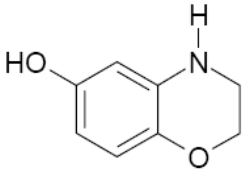
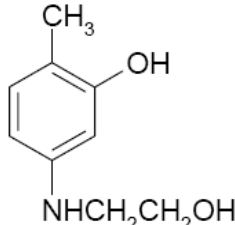
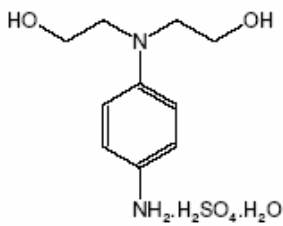
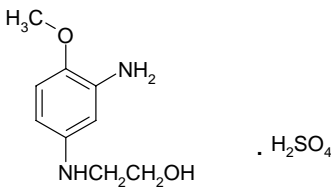
Summary

Since the nitrosation chemistry of hair dye constituents, product types, ratios and formation kinetics are difficult to predict, case by case investigations are necessary. In many cases, elucidation of structures and biological properties of the resulting NOC would represent an enormous task, since structures expected to result from nitrosation of such hair dye components are vastly diverse. Biological data on the corresponding array of generated N nitrosation products are not available. NDELA formation or contamination is only expected to result from nitrosation in very few cases. When information on a specific NOC structure is not available, the default assumption that all potentially generated NOC will be mutagenic/carcinogenic should be applied.

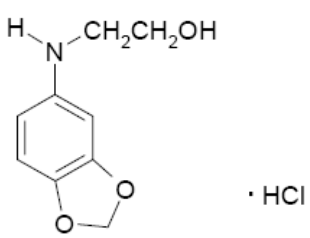
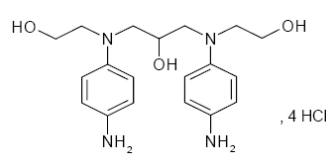
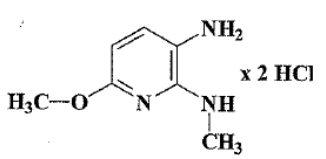
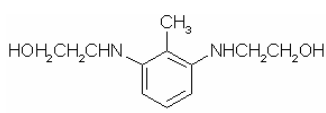
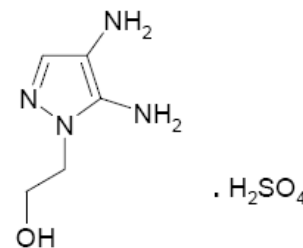
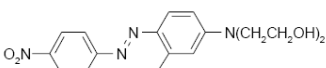
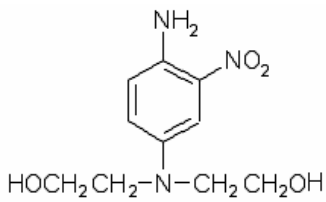
It is suggested to refer to the general purity specifications for amines in cosmetics, requiring that a given amine constituent contains not more than 50 ppb of an NOC. When compound specific analytical information is not available, generic determination of apparent total NOC (ATNC) should be used. It is suggested to follow the procedure as described in

the COLIPA 2009 guidance document. In cases where aromatic NOC are expected to be generated, N-Nitroso-N-Methylaniline may serve as a reference instead of NDELA.

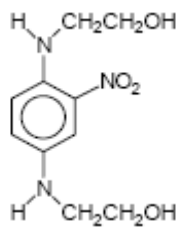
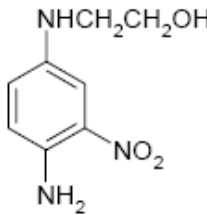
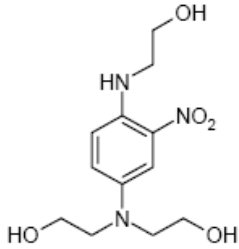
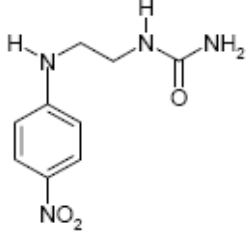
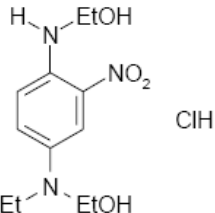
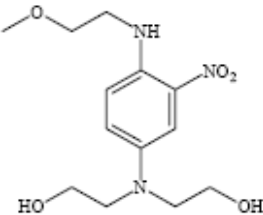
Table 1: Examples of hair dyes ingredients which may react with nitrosating agents

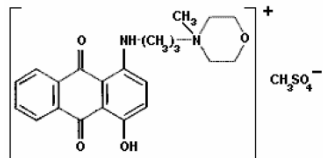
COLIPA n°	INCI name - CAS number <i>Common name in italic</i>	Formula
A9	N-Phenyl-p-phenylenediamine CAS: 101-54-2; 2198-59-6 (HCL); 4698-29-7, sulphate	
A22	p-Methylaminophenol CAS: 150-75-4	
A25	Hydroxybenzomorpholine CAS: 26021-57-8	
A31	2-Methyl-5-hydroxyethylaminophenol CAS: 55302-96-0	
A50	N,N-bis(2-Hydroxyethyl)-p-phenylenediamine sulfate CAS: 54381-16-7 (sulfate)	
A84	2-Amino-4-hydroxyethylaminoanisole CAS: 83763-47-7	

Opinion on Nitrosamines and Secondary amines in Cosmetic products

COLIPA n°	INCI name - CAS number <i>Common name in italic</i>	Formula
A98	Hydroxyethyl-3,4-methylenedioxyaniline HCl CAS: 94158-14-2	
A121	Hydroxypropyl bis (N-hydroxyethyl-p-phenylenediamine) HCl CAS: 128729-30-6 (free base); 128729-28-2 (hydrochloride)	
A130	6-Methoxy-2-methylamino-3-aminopyridine HCl CAS: 83732-72-3 (2HCl); 90817-34-8 (HCl)	
A138	2,6-Dihydroxyethylaminotoluene CAS: 149330-25-6	
A154	1-Hydroxyethyl-4,5-diamino pyrazole sulfate CAS: 155601-30-2	
B5	Disperse Red 17 (Ethanol, 2,2'-[[3-methyl-4-[(E)-(4-nitrophenyl)azo]phenyl]imino]bis-) sulfate CAS: 3179-89-3	
B31	HC Red n° 13 (2,2'-((4-Amino-3-nitrophenyl)imino)bisethanol) CAS: 94158-13-1	

Opinion on Nitrosamines and Secondary amines in Cosmetic products

COLIPA n°	INCI name - CAS number <i>Common name in italic</i>	Formula
B34	N,N'-bis(Hydroxyethyl)-2-nitro-p-phenylenediamine CAS: 84041-77-0	
B36	HC Red n° 7 (1-Amino-2-nitro-4-(β-hydroxyethyl)-aminobenzene) CAS: 24905-87-1	
B37	HC Blue n° 2 (2,2'-{[4-(2-hydroxyethyl)amino-3-nitrophenyl]imino} bisethanol) CAS: 33229-34-4	
B70	4-Nitrophenyl aminoethylurea CAS: 27080-42-8	
B73	HC Blue n° 12 (1-(β-hydroxyethyl)amino-2-nitro-4-N-ethyl-N-(β-hydroxyethyl)aminobenzene) CAS: 104516-93-0 (free base); 132885-85-9	
B77	HC Blue n° 11 (Ethanol, 2,2'-[[4-[(2-methoxyethyl)amino]-3-nitrophenyl]imino]bis-) CAS: 23920-15-2	

COLIPA n°	INCI name - CAS number <i>Common name in italic</i>	Formula
C117	Hydroxyanthroquinone aminopropyl methyl morpholinium methosulphate CAS: 38866-20-5	

4.4. Factors that influence the formation of nitrosamines

4.4.1. Nitrosation in the presence of nitrogen oxides

Is N-nitrosation expected to occur, for instance in the presence of nitrogen oxides (NOx), when pH is close to neutrality or even alkaline?

The nitrosation of amines by NOx under neutral to alkaline conditions to produce nitrosamines has been well established, both in the laboratory and in practice (Williams, 2004; Challis et al, 1977, 1978a, 1978b, 1979a, 1979b, 1979c, 1980, 1982a, 1982b; Eisenbrand et al, 1991). A review relevant to cosmetics and personal care items is provided in the COLIPA technical document (Section 4.1) (COLIPA, 2009). Major sources of NOx include: air pollution; tobacco smoke, petroleum-or gas-powered machine exhaust, nitrite treated containers; and high temperature burners, such as used for heating buildings.

4.4.2. Role of nitrites as nitrosating agents

Ionic nitrite: It is important to realize that nitrite can be formed from NOx under neutral to basic conditions (Williams, 2004), as well as by the bacterial or chemical reduction of nitrate. Ionic nitrite can react with amides of alkanolamines to produce nitrosamines as discussed in section 4.4.3 below.

Ionic nitrite reacts with aldehydes (particularly formaldehyde) and amines to produce nitrosamines through the intermediacy of imines (Williams, 2004; Keefer et al, 1973). Nitrous esters (alkyl nitrites) can also be formed by displacement reactions involving ionic nitrite from other esters (Loeppky et al, 1982, 1984). Nitrous esters are very effective nitrosating agents (Williams, 2004).

4.4.3. Risks of nitrosation during production/formulation

Amides of dialkanolamines react readily with ionic nitrite at temperatures which can be achieved from heats of mixing (Loeppky et al, 1982, 1984). Other amides in the presence of high-boiling alcohols, such as glycerol, also can react with nitrite to produce nitrosamines when heated. In the first case, the amide of the dialkanolamine, for example, a fatty acid amide, slowly rearranges both inter- and intramolecularly by acyl transfer from the nitrogen to the alcohol oxygen. Although this process involves an equilibrium that greatly favours the amide, the resulting ester rapidly reacts with ionic nitrite to form a nitrous ester which then reacts with the liberated secondary amine to form a nitrosamine. The lauryl amide of DEA forms NDELA in significant yields, when heated with nitrite. The quantities formed are increased when ethylene glycol is added.

Other tertiary amides have been shown to produce nitrosamines when heated with nitrite in a high-boiling hydroxylic solvent, such as glycerol or ethylene glycol (Loeppky et al, 1982, 1984). This process again involves acyl transfer from the amide to the alcohol to generate a

nitrous ester which then rapidly reacts with the amine produced by the acyl transfer to give a nitrosamine.

4.4.4. Role of preservatives

Some preservatives such as Bronopol (INCI 2-Bromo-2-nitropropane-1,3-diol) and Bronidox (5-Bromo-5-nitro-1,3-dioxan), permitted to be used in cosmetics, have been found to be effective nitrosating agents. In aqueous alkaline solution diethanolamine and, to a lesser extent triethanolamine (20 mmol each) yielded NDELA in substantial amounts after prolonged storage in the presence of Bronopol (80 nmol). Bronidox was a less effective nitrosating agent under the same conditions. A similar ratio in nitrosating potency was observed in a near neutral (pH 6.8) non ionic model emulsion, containing 1% diethanolamine after 50 days of storage Bronopol decomposes into nitrosating agent(s) and formaldehyde and this may contribute to enhance nitrosamine formation by the catalytic effect of formaldehyde. Thus, whenever these preservatives are coformulated with constituents having secondary amine structures, substantial nitrosamine formation is to be expected. Tertiary amines are considerably less reactive, yet can give rise to some nitrosamine formation after extended periods of storage. It is not known to what extent alkanolamides or substituted ureas might react but since their nitrosation by Bronopol/Bronidox cannot be ruled out, any co-formulation of these preservatives in cosmetics together with compounds having a nitrosatable nitrogen is to be avoided.

4.4.5. The role of catalysts like formaldehyde for nitrosation

Aldehydes (especially formaldehyde) may catalyze the formation of nitrosamines from secondary amines through the intermediacy of imines. These transformations occur readily at pH values ranging from 5-10. A similar reaction can occur with these ingredients in organic solvents that can serve as a source of formaldehyde, such as dichloromethane. The same applies also to formaldehyde donors like hydroxymethylurea and other formaldehyde releasing preservatives. Hydroxymethylsarcosine is a formaldehyde donor that may be expected to catalyse (its own) nitrosation. Thus, in addition to providing the nitrosation catalyst formaldehyde, so-called formaldehyde donor biocides can also be nitrosated as such to produce N-nitroso compounds. Practical methods for preventing such nitrosation chemistry are discussed in the Colipa document (COLIPA, 2009).

Halide ions, thiosulfate, thiols, and some phenolic compounds are all effective catalysts of nitrosation (Williams, 2004). Nitrosation can also be catalyzed by a number of other substances (see reference 1 for a complete list).

Alkanolamines are known to react with aldehydes to form oxazolidines, which then may react with a nitrosating agent to give N-nitrosooxazolidines. Both 1,3-N-nitrosooxazolidine and 5-methyl-1,3-N-nitrosooxazolidine have been found in commercial formulations containing ethanolamine (Ellen et al, 1982) and 1-amino-2-propanol (Stephany et al, 1978), respectively. These transformations also involve formaldehyde. Such heterocyclic nitrosamines could form if their precursor ingredients are present or generated in formulations of the products being considered here.

4.4.6. Inhibitors of nitrosation

Cosmetic formulations require both hydrophobic and hydrophilic nitrosation inhibitors to be effective in both phases. Water soluble inhibitors may comprise ascorbic acid/ascorbates and other water soluble antioxidants, oil soluble inhibitors may include ascorbyl palmitate, tocopherols, butylated hydroxytoluene / hydroxyanisole (BHT/BHA), gallate esters, amongst

others. There are several general reviews of methods for inhibiting the formation of nitrosamines under various circumstances (Eisenbrand 1990, Mirvish, 1975a, 1975b; Conrad et al, 1990; Loepky et al, 1994b). Annex 1 of the Colipa document (COLIPA, 2009) addresses this subject, listing a number of inhibitors.

4.4.7. Mitigation of nitrosation in raw materials and during production

The use of raw materials complying with purity specifications is required, especially with respect to contents of secondary amine, and to avoiding any contact with adventitious nitrosating agents such as nitrite treated raw material containers, atmospheric NOX-sources at production, packaging and storage units and further mitigation measures as discussed before and in the COLIPA document (COLIPA, 2009).

4.4.8. Definition of nitrosating agent/system

Nitrosating agents have been covered under 4.4.1. to 4.4.4. Any situation where nitrosating agents are present simultaneously with N nitrosatable structures, irrespective of absence or presence of catalysts and/or inhibitors, may be viewed as a nitrosating system as mentioned in the present regulation. Nevertheless, the SCCS recommends using the term nitrosating agent instead of nitrosating system in the regulation.

4.5. Carcinogenic potency of nitrosamines in cosmetics/ranking/ dose descriptors

List the nitrosamines found in cosmetics and advise the Commission of approaches to rank nitrosamines that may occur in cosmetics with respect to their carcinogenic potency.

4.5.1. List of nitrosamines in cosmetics

Nitrosamines found in cosmetics are described below. Mainly NDELA and NBHPA have been found in cosmetics. The other nitrosamines of the list have been found only very rarely in cosmetics and/or raw materials. Recent analytical surveys, covering the time period from 2000 to present also confirm that these two nitrosamines were exclusively found (BLGL, 2010; CVUA, 2005, 2006, 2007, 2008; Eisenbrand et al, 1991; RIVM, 2008, 2009; Schothorst et al, 2001, 2005).

N-nitrosated sunscreen agents, such as the NOC of 2-ethylhexyl-p-N,N-dimethylamino benzoate have been reported to occur in sunscreen formulations from the US market.

Table 2: List of Nitrosamines ever found in cosmetics and/or raw materials used for cosmetics with CAS Number and the most commonly used abbreviation when available

NDMA	N-Nitroso-dimethylamine,	CAS 62-75-9
NDEA	N-Nitroso--diethylamine,	CAS 55-18-5
	N-Nitroso-N-methyl-N-dodecylamine,	CAS 55090-44-3
NMOR	N-Nitroso- morpholine,	CAS 59-89-2
NPYR	N-Nitroso- pyrrolidine,	CAS 930-55-2
NBHPA	N-Nitroso-bis (2-hydroxypropylamine),	CAS 53609-64-6
NDELA	N-Nitroso- diethanolamine.	CAS 116-54-7
NPABA	N-nitroso-Para amino benzoic acid esters	

4.5.2. Ranking with respect to carcinogenic potential

Calculation of dose descriptors

The SCCS based the ranking of the NOCs listed in Table 2 with respect to their carcinogenic potential using dose descriptors common in carcinogenic risk assessment, noting that other approaches to rank carcinogenic potency may be considered as well.

The T25

The T25 approach is defined as the chronic dose rate (usually expressed in units of mg per kg bodyweight per day) which will give tumours at a specific tissue site in 25% of the animals after correction for spontaneous incidence and within the standard life time of the species (Dybing et al., 1997). The T25 values are likely to be within the range of the experimental data. The use of data from the lowest dose giving a significant response should in most instances reduce the problem of intercurrent mortality to an acceptable degree. This often occurs in the two years cancer bioassay when e.g. calculating the TD50 value. The data profile needed for calculating a T25 value has to be less specific, e.g. time to tumour data are not needed. It is recognized that the potential loss of precision does not match the order of magnitude differences in carcinogenic potencies found between high and low potency substances in animals.

In a study of 110 substances an almost perfect (slope in log-log plot of 1.05) coincidence was found between potency estimates by the TD50 approach and the T25 method (Dybing et al., 1997). It was concluded that, given the very large variation in carcinogenic potency between individual carcinogens, any difference between the T25 value and a "true" potency value should be negligible.

The T25 method (Dybing et al., 1997, Sanner et al., 2001) has been used within EU in setting specific concentration limits for carcinogens in preparations (EC 1999) and recently as a basis for calculation of lifetime cancer risk (LCR) and for quantitative hazard assessment of non-threshold carcinogens in several regulatory areas e.g. ECHA (2008), SCHER/SCCP/SCENIHR (2009) and SCCS (2010).

The Benchmark Dose (BMD)

The BMD approach has been increasingly used and recommended (EFSA, 2009, 2011; SCHER/SCCP/SCENIHR (2009); SCCS's Notes of Guidance (2010)). It is the dose level, derived from the estimated dose-response curve, associated with the specific change in the response defined through that BMR. It uses all available dose-response data of a study and fits a set of mathematical models. The lower one-sided confidence bound BMDL accounts for the statistical uncertainty in the data (with the statistical certainty level of 95%) and is used as a point of departure (PoD), the reference point (RP) by EFSA. The BMDL is calculated for a specified Benchmark response level (BMR). A BMR=10% of extra risk over background has been set as default, (EFSA 2009) level when analyzing cancer bioassays.

The TD₅₀

The TD₅₀ value has been introduced primarily for ranking of carcinogens in the Carcinogenic Potency Data Base (CPDB), not for the risk assessment and possible extrapolation to low doses. It is defined as dose in mg/kg bw/day which, if administered chronically for the standard lifespan of the species, will halve the probability of remaining tumourless throughout that period. Its definition requires linking with the statistical analysis method applied on individual time-to-tumour data. The ranking is based on a central tendency parameter, the median time to tumour (50% of the animals remaining tumour free),

adjusted for the respective median in the control group. In principle the TD-50 is only applicable when individual time-to-tumour data are available and can be analysed by appropriate statistical methods, e.g. excluding competing risks and intercurrent mortality. In all other cases the determination of TD-50 values is complicated by intercurrent deaths due to causes other than tumorigenesis and the non-observability of the time of tumour onset. This requires assumptions on the nature of the data and the type of dose-response. For instance, if an animal died bearing tumour(s), the time of death becomes the time of tumour onset. Consequently, the measure of tumour incidence is confounded with mortality and biased TD50 estimates can occur (Portier et al, 1987). If, on the other hand, tumours do not significantly alter survival, then TD50 values become related to the rate-of-death-with-tumour, rather than the tumour incidence rate (Meier et al., 1993). The TD-50 has been introduced as non-parametric statistical approach and thus not model dependent. However, it can also be used with parametric time-to-tumour models (e.g. the Weibull model).

Other Approaches

The Slope Factor (SF)

The slope factor has been used by US EPA as convenient descriptor of cancer potency (<http://www.epa.gov/iris/carcino.htm>) characterizing the slope of the dose-response curve at low doses (where the slope is still linear). It has dimensions of risk of cancer per unit dose and converts estimated daily intakes averaged over a lifetime of exposure directly to incremental risk of an individual developing cancer. The SF approach is performed in three steps: (1) selection of the appropriate data sets to use, (2) derivation of estimates at low doses from experimental data at high doses, using an extrapolation model and (3) choice of an equivalent human dose when animal data sets are used. Since the SF characterizes the slope of the dose-response curve at low doses by a linear approximation it can be considered as a specific modification of one model class (the linearized multistage model) also used in the BMD approach.

Quantitative structure toxicity relationship (QSTR) information may be used to qualitatively define carcinogenic potencies e.g., by establishing several categories of potency. TOPKAT(®) has been used to identify chemicals with a high probability of being chronically toxic and/or carcinogenic (Bull et al, 2011). Another approach to QSAR modelling is TOPological Substructural MOlecular DEsign (TOPS-MODE) for predictions of the carcinogenic potency of nitroso compounds (Helguera et al, 2010).

Dose-response data in various *in vivo* genotoxicity studies. The *in vivo* genotoxicity potency (e.g. micronucleus assay (MN), the *in vivo* transgenic rodent mutation assay (TG), chromosome aberration and the comet assay) of different substances measured as the lowest effective dose giving response or using the BMD approach have been compared with the respective T25 or BMD10 from carcinogenicity studies in rats and mice. A good correlation was found between the *in vivo* genotoxic potency and the carcinogenic potency. Thus the *in vivo* genotoxicity potency may be used as surrogate for carcinogenic potency, provided the *in vivo* genotoxicity dose-response data are of sufficient quality (Sanner et al., 2005, Hernandez et al., 2011)

The SCCS decided to use the three dose descriptors T25, BMD and TD50 for ranking carcinogenic potency. Comparative ranking is summarized in Table 3 (columns 2-4)

Table 3: Ranking of potencies of nitrosamines based on carcinogenicity studies on rats (the data are taken from the calculations presented in Tables 1 – 7, see annexes I and II)

Name	T25 (mg/kg bw/d) (SD)	BMDL-10 (mg/kg bw/d)	TD50 (mg/kg bw/d)
N-nitrosodimethylamine NDMA	0.058 (±0.028)	0.027	0.0959 ^{m,v}
N-nitrosodiethylamine NDEA	0.085 (±0.065)	0.018 #	0.0265 ^{m,v}
N-nitrosomorpholine MMOR	0.094 (±0.036)	0.7 ^{&}	0.109 ^m
N-nitroso-N-methyl-N-dodecylamine	0.46 (±0.08)	NA	0.537 ^{m,P}
N-nitrosobis(2-hydroxypropyl)amine NBHPA	0.54	NA	0.846 ^m
N-nitrosopyrrolidine NPYR	0.57 (±0.46)	0.16	0.799 ^{m,P}
N-nitrosodiethanolamine NDELA	2.09 (±0.78)	0.73 [§]	3.17 ^{m,v}
N-nitroso-Para amino benzoic acid esters NPABA	No data	No data	No data

NA: not available

^m there is more than one positive experiment

^v variation is greater than ten-fold among statistical significant (two-tailed p<0.1) TD50 values from different positive experiments

^P 100% of dosed animals had tumours at a target site in an experiment in this species)

^tTD50 based on the same data as T25 was 0.161 mg/kg bw/d.

BMDL-10 value 0.034 of the study of Berger et al. (1987, 1990) was by factor of about 2 higher

§ the BMDL-value of 1.74 of the study of Berger et al. (1987, 1990) was by a factor of three higher. However, when combining the two studies of Preussmann et al. and Berger et al. a BMDL-10 = 0.73 resulted which supported this value.

In order to perform full BMD analysis, at least 3 dose groups should be available. This limits the number of substances/studies for the calculation of BMD. For the calculation of T25 only the lowest dose giving a significant increase in tumour frequency is used. It should be noted that in most cases, studies where the lowest tumour frequency was higher than 80% or the study period was less than 40 weeks were excluded.

T25 is calculated as described by Dybing et al (1997) and is based on experimental data on rats only as such data are available for all the nitrosamines considered. Suitable mice data are only available in the case of N-nitrosodimethylamine and N-nitrosopyrrolidine. In the case of N-nitrosodimethylamine the TD50 indicated less potency in mice than in rats, while in the case of N-nitrosopyrrolidine the potency appeared to be similar in both species. In most cases, the T25 has been calculated on the basis of malignant liver tumours. T25 values were determined for all studies available for seven of the above listed eight nitrosamines (no data were available for NPABA) and when more than one study was available a mean T25 with standard deviation (SD) was reported. Details on the data used and the derivation of the T25 are given in Annex I.

In contrast to the dose descriptor based on the T25 and TD50 which is an average of all data sets suitable where at least on control and one dose group was reported, the BMD approach reported the BMDL10 values of only those data sets where sufficient dose-response data were available to apply this method such that a set of models can be fitted to the dose-response data. The absolute minimum number of experimental groups (e.g. one control and two dose groups) to apply the BMD technically is three. However, in this case major model with three parameters would not provide a reliable BMDL value. Therefore, in order to account for model uncertainty, only data sets with three dose groups and a control

were subject to a BMD analysis (see Annex II) using BMDS 2.1.2 of <http://epa.gov/NCEA/bmds/>The smallest BMDL10 value among all acceptable models fitted was identified and used to characterize the BMDL10 of a study as long as the BMD and the BMDL did not differ by more than one order of magnitude which would indicate undue extrapolation. For a specific nitrosamine the smallest BMDL among all studies amenable to a BMD analysis was chosen when the study data were of comparable quality. Otherwise the BMDL of the study of highest quality was chosen. Details of all BMD analyses used to derive the dose descriptors in Table 8 are provided in an Annex.

Comments on the ranking

Based on T25 N-nitrosodimethylamine, N-nitrosodiethylamine, and N-nitrosomorpholine were the most potent nitrosamines with T25 values ranging from 0.058 to 0.094. N-nitroso-N-methyl-N-dodecylamine, N-nitroso(2-hydroxypropyl)amine, and N-nitrosopyrrolidine showed a similar potency with T25 ranging between 0.46 and 0,57 and were on the average 6 – 7 times less potent. Only N-nitrosodiethanolamine with a T25 =2.09 was about 25 times less potent than the most potent carcinogen amongst nitrosamines.

Based on the BMDL10 NDEA appeared as the most potent carcinogen (BMDL10 ranging between 0.018 and 0.034 mg/kg bw/day based on two independent studies. Slightly less potent was NDMA with a BMDL10=0.027 mg/kg bw/day. Less potent by a factor of about 10 compared to NDEA was NPYR with a BMDL10 = 0.16 mg/kg bw/day NDELA was by a factor of about 20-35 less potent with BMDL10=0,73 mg/kg bw/day whereas for the other four nitrosamines no BMD/L value could be calculated. It must be noted that differences between the studies and the dose descriptors calculated using their data are not only due to different designs but also different types of carcinogenic endpoints. For example the BMDL10 of NPYR is based on total liver tumour incidence, whereas the BMDL10 of NDMA is based on fatal liver neoplasm. Such type of uncertainty should be taken into account when interpreting differences in ranking.

The ranking by TD50 exhibited NDEA as most potent carcinogen (TD50= 0.03 mg/kg/day) whereas NDMA and NMOR were less potent (TD50 = 0.10 and 0.11 mg/kg/day, respectively. N-nitroso-N-methyl-N-dodecylamine, NPYR and showed similar potencies (TD50 = 0.54 -0.84 mg/kg/day) whereas NDELA was less potent with a TD50 = 3.2 mg/kg/day. It should be noted that for NDMA, NDEA, and NDELA the variation was greater than ten-fold among studies used for calculation of the numbers given in Table 3.

As evident from Table 3, the most commonly found nitrosamine in cosmetics, NDELA, is much less potent as compared to NDMA or NPYR, nitrosamines that are commonly found in food.

4.6. Carcinogenic potential of nitrosamines in relation to their parent compounds

There is no obvious way to relate carcinogenic potential of nitrosamines formed with the parent chemical group. To predict the carcinogenicity potency of nitrosamines potentially generated, it has to be considered they have to undergo metabolic activation, in most cases by CYP-450 dependent hydroxylation in the alfa-position, to become (geno)-toxic agents. The metabolic introduction of the hydroxyl group leads to a very unstable alfa-hydroxy-nitrosamine that decays by aldehyde elimination, liberating the ultimate electrophilic, genotoxic/mutagenic/ intermediate that subsequently binds to DNA, thereby causing mutations and cancer. Structural elements that inhibit this CYP dependent metabolic activation/electrophil generation may decrease or even abrogate mutagenicity/carcinogenicity. As an example an alpha-t-butyl substituent adjacent to the N-Nitroso group, eliminates carcinogenicity, as demonstrated e.g. for ethyl-t-butyl nitrosamine. In addition, strongly basic or acidic centers present in a given NOC may also

strongly reduce or even abrogate mutagenicity/carcinogenicity as shown e.g. for N-nitroso-N' methyl-piperazine or for most N-Nitroso-amino acids and their esters (Preussmann, 1990).

These limited structure activity observations indicate specific structural elements, when present in secondary amines may result in generation of non-carcinogenic nitrosamines. There might be other structures that can mitigate formation of carcinogenic NOC but this need to be demonstrated. It is recommended that case by case evaluations should be performed if this principle of using "safe amines" in cosmetics is to be applied (Janzowski et al, 2000).

SCCS recommends for situations where contact with nitrosating agents during production, formulation, storage and use cannot reliably be excluded, to use amines for cosmetics that are not easily nitrosated and /or give rise to non-carcinogenic nitrosamines.

4.7. Levels of nitrosamines in finished products and raw materials

Comment on the levels of 50 µg nitrosamine/ kg as set out currently in the Annexes of Directive 76/768/EEC. Should it apply to finished products or to raw materials? Should it be considered for all nitrosamines potentially formed? Should it be modified, following the ranking of carcinogenic potency of nitrosamines in question? Comment on the "maximum secondary amine content (5% in raw materials and 0.5% in finished products) ".

As currently regulated, the purity specification of 50 µg nitrosamine/ kg should apply to raw materials and to all nitrosamines potentially formed. This limit (50 µg nitrosamine/ kg) does not apply to finished products. The secondary amine content in the finished product determines the content of nitrosamines potentially formed. The secondary amine content in raw materials should therefore be as low as achievable, following GMP rules, but should not exceed the limits as laid down in the Directive. The SCCS supports the present regulation since it provides a high degree of consumer protection.

4.8. Specific cases

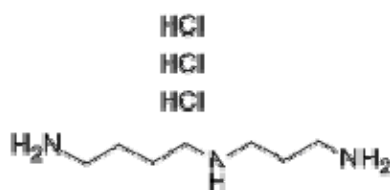
On the basis of the answers above SCCS to pronounce itself

4.8.1- on the specific cases of spermidine (CAS 334-50-9), gerotine (CAS 71-44-3) and dipropylenetriamine (CAS 56-18-8);

4.8.2- on the "Maximum secondary amine content: 5% (applies to raw materials)" and that "Maximum secondary amine content: 0.5%" in the finished cosmetic products" for the Fatty acid dialkylamides and dialkanolamines listed in entry 60 of Annex III, part I.

4.8.1 Specific Cases

The SCCS is of the opinion that the specific cases of spermidine (CAS 334-50-9), gerotine (CAS 71-44-3) and dipropylenetriamine (CAS 56-18-8) fall under the term secondary alkylamine since each of these compounds contains at least one secondary amino group, with two alkyl substitutions.



spermidine trihydrochloride, (CAS 334-50-9)



Spermine, (CAS 71-44-3), Gerotine



dipropylenetriamine (CAS 56-18-8)

It has been shown that the acidic nitrosation of spermidine produces nitrosamines that are mutagenic (Bills et al, 1973; Hildrum et al, 1975, 1976, 1977a, 1977b, Hotchkiss, 1977, 1979). The secondary amine is nitrosated in all of the nitrosamine products found. The primary amine groups are nitrosated and deaminated through the formation of short-lived reactive diazonium ions. It is reasonable to expect that other polyamines, will behave similarly. For example, some other nitrogen containing compounds, as exemplified by structures like chlorohexidine (1,1'-Hexamethylenebis[5-(4-chlorophenyl)biguanide, hexidine]) or hexetidine (1,3-Bis(2-ethylhexyl)-hexahydro-5-methylpyrimidine-5-amine,) can also be rapidly nitrosated to form nitrosamines. An excess of nitrous acid was used in that work Bae et al, 1994; Loepky et al, 1994a). Limited quantities of nitrosating agents may produce somewhat different results but it is highly likely that the secondary amine functionality will be nitrosated to give a nitrosamine. Subsequent research has verified the mutagenicity of nitrosospermidine (Sedgwick, 1997) and demonstrated that these compounds may occur in the human environment (Mah et al, 2005; Perera, 2006).

4.8.2 Secondary amine content in raw material and in finished cosmetic products

As described above (see Section 4.7), the SCCS supports the present regulation on the "*Maximum secondary amine content: 5% (applies to raw materials)*" and that "*Maximum secondary amine content: 0.5% in the finished cosmetic products*" for the fatty acid dialkylamides and dialkanolamines listed in entry 60 of Annex III, part I.

5. ABBREVIATIONS

ATNC	Apparent Total N-nitroso group Content
BMD	Benchmark dose
BMDL	Benchmark dose lower 95% confidence limit
CAS	Chemical Abstracts Service
Colipa	the European Cosmetics Association
CPDB	Carcinogenic Potency Database
DEA	Diethanolamine
NDELA	N-nitrosodiethanolamine
NO	nitrogen oxide
NOC	N-Nitroso compounds
QSAR	Quantitative structure Activity relationship
SCCNFP	Scientific Committee on Cosmetic Products and Non-Food Products intended for Consumers
SCCS	Scientific Committee on Consumer Safety

T25	The dose giving a 25% increase in the frequency of a specific tumour during the standard lifetime of the species studied.
TD50	Median Toxic Dose
WHO	World Health Organization

6. REFERENCES

- Bae, J.-Y. et al., (1994) The Nitrosation of Hexetidine and Hexedine: Characterization of the Major Nitrosamine from Common Antimicrobial Agents. *Chem. Res. Toxicol.*, 1994. 7(6): p. 868-876.
- Bayerisches Landesamt für Gesundheit und Lebensmittelsicherheit (BLGL), Jahresberichte 2007-2010, <http://www.lgl.bayern.de/publikationen/jahresberichte.htm>
- Berger, M.R. et al. (1987) Combination experiments with very low doses of three genotoxic N-nitrosamines with similar organotropic carcinogenicity in rats. *Carcinogenesis* 8: 1635-1643.
- Berger, M.R. et al. (1990). Implications of the carcinogenic hazard of low doses of three hepatocarcinogenic N-nitrosamines. *Jpn J Cancer Res* 81: 598-606.
- Bills, D.D. et al. (1973a) Potential precursors of N-nitrosopyrrolidine in bacon and other fried foods. *Journal of Agricultural and Food Chemistry*, 1973. 21(5): 876-7.
- Bull, R.J. et al. (2011) Potential carcinogenic hazards of non-regulated disinfection by-products: Haloquinones, halo-cyclopentene and cyclohexene derivatives, N-halamines, halonitriles, and heterocyclic amines. *Toxicology*. 2011 Aug 15;286(1-3):1-19. Epub 2011 May 14.
- Challis, B.C. et al. (1977) Rapid formation of carcinogenic N-nitrosamines in aqueous alkaline solutions. *Br. J. Cancer*, 1977. 35(5): p. 693-6.
- Challis, B.C. et al. (1978a) Rapid formation of N-nitrosamines from nitrogen oxides under neutral and alkaline conditions. *IARC Sci. Publ.*, 1978. 19(Environ. Aspects N-Nitroso Compd.): p. 127-42.
- Challis, B.C. et al. (1978b) The chemistry of nitroso compounds. Part 12. The mechanism of nitrosation and nitration of aqueous piperidine by gaseous dinitrogen tetroxide and dinitrogen trioxide in aqueous alkaline solutions. Evidence for the existence of molecular isomers of dinitrogen tetroxide and dinitrogen trioxide. *J. Chem. Soc., Perkin Trans. 2*, 1978(12): p. 1296-302.
- Challis, B.C. et al. (1979a) Rapid nitrosation of amines in aqueous alkaline solutions by beta-substituted alkyl nitrites. *J. Chem. Soc., Chem. Commun.*, 1979(7): p. 315-16.
- Challis, B.C. et al. (1979b) The chemistry of nitroso compounds. Part 15. Formation of N-nitrosamines in solution from gaseous nitric oxide in the presence of iodine. *J. Chem. Soc., Perkin Trans. 1*, 1979(11): p. 2768-75.
- Challis, B.C. et al. (1979c) The chemistry of nitroso-compounds. Part 11. Nitrosation of amines by the two-phase interaction of amines in solution with gaseous oxides of nitrogen. *J. Chem. Soc., Perkin Trans. 1*, 1979(2): p. 299-304.
- Challis, B.C. et al. (1980) New pathways for the rapid formation of N-nitrosamines under neutral and alkaline conditions. *IARC Sci. Publ.*, 1980. 31(N-Nitroso Compd.: Anal., Form. Occurrence): p. 43-58.
- Challis, B.C. et al. (1982a) Formation of N-nitrosamines and N-nitramines by photolysis. *IARC Sci. Publ.*, 1982. 41(N-Nitroso Compd: Occurrence Biol. Eff.): p. 31-40.
- Challis, B.C. et al. (1982b) Formation of N-nitrosamines and N-nitramines by gaseous nitrogen dioxide. *Acta Cient. Compostelana*, 1982. 19(1-2): p. 153-66.
- Chemisches und Veterinäruntersuchungsamt (CVUA) Karlsruhe, Jahresbericht 2006, http://www.ua-bw.de/pub/beitrag.asp?subid=2&ID=991&Thema_ID=1&lang=DE
- Chemisches und Veterinäruntersuchungsamt (CVUA) Karlsruhe, Jahresbericht 2005, http://www.ua-bw.de/pub/beitrag.asp?subid=2&ID=992&Thema_ID=1&lang=DE

- Chemisches und Veterinäruntersuchungsamt (CVUA) Stuttgart, Jahresbericht 2008, <http://jahresbericht.cvuas.de/docs/jb2008.pdf>
- Chemisches und Veterinäruntersuchungsamt (CVUA) Stuttgart, Jahresbericht 2007, <http://jahresbericht.cvuas.de/docs/jb2007.pdf>
- Chemisches und Veterinäruntersuchungsamt (CVUA) Stuttgart, Jahresbericht 2006, http://cvuas.untersuchungsämter-bw.de/pdf/cvuas_jahresbericht2006.pdf
- Chemisches und Veterinäruntersuchungsamt (CVUA) Stuttgart, Jahresbericht 2005, http://cvuas.untersuchungsämter-bw.de/pdf/cvuas_jahresbericht2005.pdf
- Colipa, Technical Guidance Document on Minimising and Determining Nitrosamines in Cosmetics. 2009, Colipa, The European Cosmetics Association. p. 43.
- Conrad, J. et al (1990) Nitrosation of Bromhexine, in The Significance of N-Nitrosation of Drugs, G. Eisenbrand, G. Bozler, and H.v. Nicolai, Editors. 1990, Gustav Fischer Verlag: Stuttgart. p. 181-196.
- Dybing E. et al. (1997) T25: A simplified carcinogenic potency index: Description of the system and study of correlations between carcinogenic potency and species/site specificity and mutagenicity. *Pharmacology and Toxicology* 80, 272-279 (1997).
- ECHA (2008) Guidance on information requirements and chemical safety assessment; Chapter R.7c: Endpoint specific guidance, European Chemicals Agency, 2008
- EFSA (2009) Guidance of the Scientific Committee on Use of the benchmark dose approach in risk assessment. <http://www.efsa.europa.eu/en/efsajournal/pub/1150.htm>
- EFSA (2011) Use of BMDS and PROAST software packages by EFSA Scientific Panels and Units for applying the Benchmark Dose (BMD) approach in risk assessment. <http://www.efsa.europa.eu/en/supporting/doc/s113e.pdf>
- Eisenbrand, G. (1990) Endogenous nitrosation - findings and problems, in The Significance of N-Nitrosation of Drugs, G. Eisenbrand, G. Bozler, and H.v. Nicolai, Editors. 1990, Gustav Fischer Verlag: Stuttgart. p. 47-68.
- Eisenbrand, G. et al. (1991) N-Nitrosoalkanolamines in cosmetics. In: Relevance to human cancer of N-nitroso compounds, tobacco smoke and mycotoxins, I.K. O'Neill, J. Chen and H. Bartsch, (eds.), 1991, IARC Sci. Publ. 105, pp. 238-241.
- Ellen, G. et al. (1982) N-Nitroso-1,3-oxazolidine, a carcinogenic compound detected as an impurity in a commercial cutting fluid. *Recl.: J. R. Neth. Chem. Soc.*, 1982. 101(10): p. 357-9.
- Fitzpatrick, R.B. (2008) CPDB: Carcinogenic Potency Database. *Med Ref Serv Q.* 2008 Fall;27(3):303-11.
- Gowenlock, B. et al. (1979) Nitrosative dealkylation of some symmetrical tertiary amines. *J. Chem. Soc., Perkin Trans. 2*, 1979: p. 1110-1114.
- Helguera, A.M. et al. (2010) Quantitative structure-activity relationship modelling of the carcinogenic risk of nitroso compounds using regression analysis and the TOPS-MODE approach. *SAR QSAR Environ Res.* 2010 Apr;21(3-4):277-304.
- Hernandez, L.G. et al. (2011) Can carcinogenic potency be predicted from in vivo genotoxicity data? A meta-analysis of historical data. *Environ Mol Mutagen* 52: 518-528, 2011.
- Hildrum, K.I. et al. (1975) Identification of gamma -butenyl-(beta -propenyl)nitrosamine, the principal volatile nitrosamine formed in the nitrosation of spermidine or spermine. *J. Agric. Food Chem.*, 1975. 23(1): p. 34-7.
- Hildrum, K.I. et al. (1976) Nitrosamines from the nitrosation of spermidine and spermine. *IARC Sci Publ*, 1976(14): p. 205-14.

- Hildrum, K.I. et al. (1977a) Factors influencing the rate of formation of volatile N-nitrosamines during the nitrosation of spermidine. *J. Agric. Food Chem.*, 1977. 25(2): p. 255-7.
- Hildrum, K.I. et al. (1977b) Formation of volatile, hydroxylated, and chlorinated N-nitrosamines during the nitrosation of spermidine 3-hydrochloride. *J Agric Food Chem*, 1977. 25(2): p. 252-5.
- Hotchkiss, J.H. et al. (1977) Formation of bis(hydroxyalkyl)-N-nitrosamines as products of the nitrosation of spermidine. *J. Agric. Food Chem.*, 1977. 25(5): p. 1183-9.
- Hotchkiss, J.H. et al. (1979) Mutagenicity of nitrosamines formed from nitrosation of spermidine. *Mutat Res*, 1979. 68(3): p. 195-9.
- Huxel, E.T. et al. (1974) Formation of N-nitrosopyrrolidine from pyrrolidine ring containing compounds at elevated temperatures. *Journal of Agricultural and Food Chemistry*, 1974. 22(4): p. 698-700.
- Janzowski, C. et al. (2000) in Wichmann, Schlipkötter, Füllgraf, eds, *Handbuch Umweltmed.11/00,Org.Verb./N-Nitrosamine IV-4.1-30*, 2000
- Jones, A.R. et al. (1973) Steric effects in the nitrosation of piperidines. *Cancer Res.*, 1974. 34(5): p. 1079-81.
- Keefer, L.K. et al. (1973) N-nitrosation by nitrite ion in neutral and basic medium. *Science*, 1973. 181(4106): p. 1245-1247.
- Loeppky, R.N. et al. (1982) Nitrosamine formation from ternary nitrogen compounds. *International Agency for Research in Cancer Scientific Publication*, 1982. 41: p. 41-56.
- Loeppky, R.N. et al. (1983) Stereoelectronic Effects in Tertiary Amine Nitrosation: Nitrosative Cleavage vs. Aryl Ring Nitrosation. *J. Org. Chem.*, 1983. 48: p. 2751-2756.
- Loeppky, R.N. et al. (1984) Ester-mediated nitrosamine formation from nitrite and secondary or tertiary amines. *International Agency for Research in Cancer Scientific Publication*, 1984. 57: p. 353-363.
- Loeppky, R.N. et al. (1990) Rapid nitrosamine formation from tertiary nitrogen compounds: An overview, in *Significance of N-Nitrosation of Drugs (Drug Dev. Eval)*, G. Eisenbrand, G. Bolzer, and H.v. Nicolai, Editors. 1990, Gustav Fischer Verlag: New York.
- Loeppky, R.N. et al. (1991) Nitrosation of tertiary aromatic amines related to sunscreen ingredients, in *Relevance of Human Cancer of N-Nitroso Compounds, Tobacco Smoke, and Mycotoxins*, I.K. O'Neill, J. Chen, and H. Bartsch, Editors. 1991, International Agency for Research on Cancer: Lyon, Fr. p. 244-252.
- Loeppky, R.N. et al. (1994a) An aziridinium ion intermediate in the nitrosation of a hexetidine model. *Chem. Res. Toxicol.*, 1994. 7: p. 661-667.
- Loeppky, R.N. et al. (1994b) Blocking nitrosamine formation: Understanding the chemistry of nitrosamine formation, in *Nitrosamines and Related N-Nitroso Compounds: Chemistry and Biochemistry*, R.N. Loeppky and C.J. Michejda, Editors. 1994, American Chemical Society: Washington, DC. p. 52-65.
- Loeppky, R.N. et al. (1998) Evidence for radical cations in linked mechanisms of N,N-dialkyl aromatic amine nitration and nitrosative dealkylation. *J. Am. Chem. Soc.*, 1998. 120(21): p. 5193-5202.
- Maduagwu E.N. (1985) Nitrosation of quaternary ammonium compounds *in vivo* in the Wistar rat. *Xenobiotica* 15(12):1061-1068.
- Mah, J.-H. et al. (2005) Influence of curing and heating on formation of N-nitrosamines from biogenic amines in food model system using Korean traditional fermented fish product. *Food Sci. Biotechnol.*, 2005. 14(1): p. 168-170.

- McGregor, D. et al (2010) Guidance for the classification of carcinogens under the Globally Harmonised System of Classification and Labelling of Chemicals (GHS). *Crit Rev Toxicol.* 2010 Mar;40(3):245-85.
- Meier, K.L. et al. (1993). A measure of tumorigenic potency incorporating dose-response shape. *Biometrics.* 49: 917-926
- Mirvish, S.S. (1970) Kinetics of dimethylamine nitrosation in relation to nitrosamine carcinogenesis. *J. Nat. Cancer Inst.,* 1970. 44(3): p. 633-639.
- Mirvish, S.S. (1975a) Blocking the formation of N-nitroso compounds with ascorbic acid in vitro and in vivo. *Ann. N. Y. Acad. Sci.,* 1975: p. 258.
- Mirvish, S.S. (1975b) Formation of N-Nitroso compounds: Chemistry, kinetics and in vivo occurrence. *Toxicology and Applied Pharmacology,* 1975. 31: p. 325-351.
- Okun, J.D. et al. (1977) Kinetics of nitrosamine formation in the presence of micelle-forming surfactants. *J. Natl. Cancer Inst.* 1977,58(2): 409-11)
- Park, M. et al. (1998) The synthesis of highly reactive multifunctional □□□-epoxy- and □-acetoxy-nitrosamines. *Tetrahedron Lett.,* 1998. 39: p. 1287-1290.
- Perera, A.M.A.. (2006) N-nitrosoamines. *Chromatogr. Sci. Ser.,* 2006. 93(Chromatographic Analysis of the Environment (Third Edition)): p. 419-452.
- Portier, C.J. et al. (1987). Issues concerning the estimation of the TD50. *Risk Anal.* 7: 437-447.
- Preussmann, R. (1990) The Significance of n-Nitrosation of drugs, 1990, G. Fischer Verlag Stuttg. New York, Carcinogenicity and Structure-Activity relationships of N-Nitroso compounds, p3-17
- RIVM (2008) Rijksinstituut voor Volksgezondheid en Milieu, Nitrosaminen (NDELA) in cosmetica, Fact sheet, Maart 2008
- RIVM (2009) Rijksinstituut voor Volksgezondheid en Milieu, Marktkennis Cosmetica, Nitrosaminen (NDELA) en Zware Metalen in Decoratieve Cosmetica, Fact Sheet, September 2009
- Sanner, T. et al (2001) A simple method for quantitative risk assessment of non-threshold carcinogens based on the dose descriptor T25. *Pharmacol Toxicol* 88: 331-341 (2001).
- Sanner, T. et al (2005) Comparison of carcinogenic and in vivo genotoxic potency estimates. *Basic Clinical Pharm Toxicol,* 96: 131-139, 2005.
- SCCS (2010) Notes of Guidance for the testing of Cosmetic Ingredients and their Safety evaluation, of 14 December 2010
- SCHER/SCCP/SCENIHR (2009) Scientific opinion on risk assessment methodologies and approaches for genotoxic and carcinogenic substances, adopted on the 19th plenary meeting of the SCCP of 21 January 2009.
- Schothorst, R.C. et al. (2001) Occurrence of N-nitrosodiethanolamine (NDELA) in cosmetics from the Dutch market. *Int. J. Cosmet. Sci.* 2001, 23, 109-114.
- Schothorst, R.C. et al. (2005) Determination of N-nitrosodiethanolamine in cosmetic products by LC-MS-MS. *Anal. Bioanal. Chem.* 2005, 381, 681-685.
- Sedgwick, B. (1997) Nitrosated peptides and polyamines as endogenous mutagens in O6-alkylguanine-DNA alkyltransferase deficient cells. *Carcinogenesis,* 1997. 18(8): p. 1561-1567.
- Smith, P.A.S. et al. (1967) Nitrosative cleavage of tertiary amines. *J. Am. Chem. Soc.,* 1967. 89(5): p. 1147-1157.

- Stephany, R.W. et al. (1978) N-Nitroso-5-methyl-1,3-oxazolidine identified as an impurity in a commercial cutting fluid. *Recueil des Travaux Chimiques des Pays-Ba*, 1978. 97(6): p. 177-8.
- Teuten, E.L. et al. (2005) The mechanistic origin of regiochemical changes in the nitrosative N-dealkylation of N,N-dialkyl aromatic amines. *Organic & Biomolecular Chemistry*, 2005. 3(6): p. 1097-1108.
- Williams, D.L.H. (2004) *Nitrosation reactions in the chemistry of nitric oxide*. 2004, Amsterdam: Elsevier. 268.

Annex I

Tables of T25 for the different compounds:

The calculations below are based on experimental data on rats only as such data are available for all the nitrosamines considered. Suitable mice data are only available in the case of N-nitrosodimethylamine and N-nitrosopyrrolidine. In the case of N-nitrosodimethylamine the TD50 was less potent in mice than in rats, while in the case of N-nitrosopyrrolidine the potency appeared to be similar. In most cases, the T25 has been calculated on the bases of malignant liver tumours.

The T25 is calculated from the experimental data as described by Dybing et al. (1997). HT25 represent the human-equivalent dose estimated from the rat studies by using a scaling factor based on body weights to the power of $\frac{3}{4}$ as described by Sanner et al. (2001). Unless the specific weight is given, the following defaults are used: humans = 70 kg, male rats 500 g and female rats 350 g (ECHA, 2008).

N-nitrosodiethanolamine (1116-54-7) NDELA

Table 1: N-nitrosodiethanolamine (CAS 1116-54-7). Calculation of T25 and HT25 from carcinogenicity studies on rats¹

T25: HT25 (mg/kg bw/day)	Target	Species	Route	Expo-sure (week)	Observation (week)	Lowest tumour frequency %	Ref.
2.19 0.64	Liver carcinomas	F344 rats male	Drinking water	100	104	20	Lijinsky and Kovatch, 1985
1.79 0.47	Liver carcinomas	F344 rats female	Drinking water	100	104	35	Lijinsky and Kovatch, 1985
2.22 0.65	Liver carcinomas	SD rats male	Drinking water	116	116	60	Preussm an et al., 1982
8.09 2.28	Liver carcinomas	SD rats male	Drinking water	130	130	1.9	Berger et al., 1990
3.21 0.94	Liver carcinomas	SD rats male	Drinking water	104	104	33	Zerban et al., 1988
1.05 0.28	Liver unspecified	SD rats female	Drinking water	50	104	70	Hecht et al., 1989

¹ Studies with exposure time less than 50 weeks and studies where the frequency of hepatocellular tumours at the lowest dose tested were 80% or higher were excluded,

The experiment by Berger *et al.*(1990) has been excluded from the calculation of the mean as it is more than 7 times standard deviation above the mean

Mean T25 = 2.09± 0.78 mg/kg bw/d (range 1.05 - 3.21 mg/kg bw/d)

Mean HT25 = 0.60 ± 0.24mg/kg bw/d (range 0.28 – 0.94 mg/kg bw/d)

TD50 = 3.17 (rat, there is more than one positive experiment, variation is greater than ten-fold among statistical significant (two-tailed p<0.1) TD50 values from different positive experiments).

N-nitrosobis(2-hydroxypropyl)amine (53609-64-6), NBHPA

Table 2. N-nitrosobis(2-hydroxypropyl)amine (Cas no. 53609-64-6). Calculation of T25 and HT25 from carcinogenicity studies on rats

T25: HT25 (mg/kg bw/day)	Target	Species	Route	Exposure (week)	Observation (week)	Lowest tumour frequency %	References
0.56 0.15	Esophagus carcinomas	F344 rats female	Drinking water	42	42	65	Lijinsky et al., 1984
0.52 0.15	Lung adenomas	Wistar rats male	Drinking water	52	52	60	Konishi et al., 1978

T25 = 0.54 ± 0.03 mg/kg bw/d (range 0.52 – 0.56 mg/kg bw/d)

HT25 = 0.15 mg/kg bw/d.

TD50 = 0.846

N-nitrosodimethylamine (62-75-9), NDMA

Table 3: N-nitrosodimethylamine (CAS 62-75-9). Calculation of T25 and HT25 from carcinogenicity studies on rats

T25: HT25 (mg/kg bw/day)	Target	Species	Route	Exposure (week)	Observation (week)	Lowest tumour frequency %	References
0.11 0.032	Lung scc	F344 rats male	Gavage	30	45	32	Lijinsky et al., 1987
0.041 0.012	Liver carcinomas	F344 rats male	Drinking water	30	80	50	Lijinsky and Reuber, 1984
0.032 0.0085	Liver carcinomas	F344 rats female	Drinking water	30	90	45	Lijinsky and Reuber, 1984
0.73 0.22	Leydig cell	Wistar rats male	Feed	54	69	47	Terao et al., (1978)
0.059 0.017	Liver malignant	Wistar rats male	Drinking water	104	104	22	Peto et al., 1991
0.044 0.012	Liver malignant	Wistar rats female	Drinking water	104	104	45	Peto et al., 1991
0.061 0.016	Liver carcinomas	Wistar rats female	Feed	93	93	18	Arai et al., 1979

Terao et al., (1978) not included in the risk assessment as it is based on Leydig cell tumours)

Peto, 1991 (cited)

“The linear relationship observed at low dose rates (below 1 ppm) suggests that under these experimental conditions, among rats allowed to live their natural life span, a dose of 1 ppm of NDEA or NDMA in the drinking water will cause about 25% to develop a liver neoplasm.”

Male rat, 25 ml per day, 500 mg; 1 ppm represent 0.001 mg/ml, 0.025 mg/0.500 = 0.050 mg/kg bw/d; female rat 20 ml per day, 350 g; 0.001 mg/ml, 0.020 mg/0.350 = 0.057

mg/kg bw/d. The above citation implies a T25 of 0.050 – 0.057 mg/kg bw/d both for NDMA and NDEA.

Mean T25 = 0.058 ± 0.028 mg/kg bw/d (range 0.032 – 0.11 mg/kg bw/d)

Mean HT25 = 0.016 ± 0.008 mg/kg bw/d (range 0.0085 – 0.032 mg/kg bw/d)

TD50 = 0.096 (rat, there is more than one positive experiment, variation is greater than ten-fold among statistical significant (two-tailed $p < 0.1$) TD50 values from different positive experiments)

N-nitrosopyrrolidine (930-55-2), NPYR

Table 4: N-nitrosopyrrolidine (CAS 930-55-2). Calculation of T25 and HT25 from carcinogenicity studies on rats¹

T25: HT25 (mg/kg bw/day)	Target	Species	Route	Exposure (week)	Observation (week)	Lowest tumour frequency %	References
0.14 0.039	Liver carcinomas	Wistar rats male + female	Drinking water	104	104	42	Gray et al., 1991
0.62 0.17	Liver carcinomas	SD rats male	Drinking water	132	132	15	Berger et al., 1987
0.31 0.090	Liver carcinomas	SD rats male	Drinking water	86	104	66	Hoos et al., 1985
1.19 0.33	Liver carcinomas	SD rats male+ female	Drinking water	104	104	21	Preussman et al., 1977
0.43 0.12	Liver carcinomas	F344 rats male	Drinking water	80	80 weeks	87	Chung et al., 1986
0.27 0.070	Liver carcinomas	F344 rats female	Drinking water	30	104	100	Michejda et al., 1986
0.44 0.12	Liver carcinomas	F344 rats female	Drinking water	50	100	100	Lijinsky, Reuber 1981
0.61 0.18	Liver tumours	F344 rats male	Drinking water	50	80	86	Lijinsky, Taylor 1976
0.81 0.22	Liver tumours	F344 rats female	Drinking water	50	75	87	Lijinsky, Taylor 1976

¹ Studies with exposure time less than 50 weeks and studies where the frequency of hepatocellular tumours at the lowest dose tested were 80% or higher were excluded,

Mean T25 = 0.57 ± 0.46 mg/kg bw/d (range 0.14 – 1.19 mg/kg bw/d)

Mean HT25 = 0.16 ± 0.13 mg/kg bw/d (range 0.039 – 0.33 mg/kg bw/d)

(Note: Including all results will not change the results.

Mean T25 = 0.54 ± 0.32 mg/kg bw/d (range 0.14 – 1.19 mg/kg bw/d)

Mean HT25 = 0.15 ± 0.09 mg/kg bw/d (range 0.039 – 0.33 mg/kg bw/d)

TD50 = 0.799 (rat, there is more than one positive experiment, 100% of dosed animals had tumours at a target site in an experiment in this species)

N-nitrosomorpholine (CAS 59-89-2), NMOR

Table 5: N-nitrosomorpholine (CAS 59-89-2). Calculation of T25 and HT25 from carcinogenicity studies on rats

T25: HT25 (mg/kg bw/day)	Target	Species	Route	Exposure (week)	Observation (week)	Lowest tumour frequency %	References
0.123 0.033	Liver carcinomas	F344 rats female	Drinking water	50	100	10	Lijinsky et al., 1988
0.047 0.014	Liver carcinomas	F344 rats male	Drinking water	50	80	90	Lijinsky, Reuber 1982
0.122 0.032	Liver carcinomas	F344 rats female	Drinking water	50	55	85	Lijinsky, Reuber, 1982
0.082 0.022	Liver carcinomas	F344 rats female	Drinking water	50	65	95	Hecht et al., 1989

Mean T25 = 0.094 ± 0.036 mg/kg bw/d (range 0.047 – 0.123 mg/kg bw/d)

Mean HT25 = 0.025 ± 0.009 mg/kg bw/d (range 0.014 – 0.033 mg/kg bw/d)

TD50 = 0.109 (rat, there is more than one positive experiment)

N-nitroso-N-methyl-N-dodecylamine (55090-44-3),

Table 6: N-nitroso-N-methyl-N-dodecylamine (CAS 55090-44-3). Calculation of T25 and HT25 from carcinogenicity studies on rats

T25: HT25 (mg/kg bw/day)	Target	Species	Route	Exposure (week)	Observation (week)	Lowest tumour frequency %	References
0.40 0.11	Bladder carcinomas	F344 rats female	Gavage	30	50	84	Lijinsky et al., 1983
0.52 0.15	Bladder carcinoma	F344 rats male	Gavage	30	104	95	Lijinsky et al., 1981

Mean T25 = 0.46 ± 0.08 mg/kg bw/d (range 0.40 – 0.52 mg/kg bw/d)

Mean HT25 = 0.13 ± 0.03 mg/kg bw/d (range 0.11 – 0.15 mg/kg bw/d)

Note: Very high tumour frequency at the lowest doses.

TD50 = 0.537 (rat, there is more than one positive experiment, 100% of dosed animals had tumours at a target site in an experiment in this species)

N-nitrosodiethylamine (55-18-5), NDEA

Table 7: N-nitrosodiethylamine (CAS 55-18-5). Calculation of T25 and HT25 from carcinogenicity studies on rats

T25: HT25 (mg/kg bw/day)	Target	Species	Route	Exposure (week)	Observation (week)	Lowest tumour frequency %	References
0.11 0.032	Liver carcinomas	F344 rats male	Gavage	30	45	32	Lijinsky et al., 1987
0.11 0.031	Liver Malignant	Wistar rats male	Drinking water	104	104	8	Peto et al., 1991
0.042 0.011	Liver Malignant	Wistar rats female	Drinking water	104	104	32	Peto et al., 1991
0.0088 0.0023	Liver carcinomas	F344 rats female	Drinking water	60	104	30	Lijinsky et al., 1981
0.19 0.054	Liver carcinomas	SD rats Male	Drinking water	122	122	26	Berger et al., 1987
0.030 0.0081	Esophagus unspecified	F344 rats female	Drinking water	30	45	89	Lijinsky et al., 1983
0.049 0.014	Liver unspecified	SD rats male	Drinking water	116	116	45	Habs, Schmähl, 1980

Mean T25 = 0.085 ± 0.065 mg/kg bw/d (range 0.0088 – 0.19 mg/kg bw/d)

Mean HT25 = 0.024 ± 0.019 mg/kg bw/d (range 0.0023 – 0.054 mg/kg bw/d)

TD50 = 0.0265 (rat, there is more than one positive experiment, variation is greater than ten-fold among statistical significant (two-tailed $p < 0.1$) TD50 values from different positive experiments) (TD50 based on the same data as T25 gives TD50 = 0.161 mg/kg bw/d).

References

Arai M, Aoki Y, Nakanishi K, Miyata Y, Mori T, Iyo N (1979). Long-term experiment of maximal non-carcinogenic dose of dimethylnitrosamine for carcinogenesis in rats. *Gann* 70: 549-558.

Berger MR, Schmähl D, Zerban H (1987). Combination experiments with very low doses of three genotoxic N-nitrosamines with similar organotropic carcinogenicity in rats. *Carcinogenesis* 8: 1635-1643.

Berger MR, Schmähl D, Edler L (1990) Implications of the carcinogenic hazard of low doses of three hepatocarcinogenic N-nitrosamines. *Jpn J Cancer Res* 81: 598-606.

Chung FL, Tanaka T, Hecht SS (1986). Induction of liver tumors in F344 rats by crotonaldehyde. *Cancer Research* 46: 1285-1289.

Gray R, Peto R, Brantom P, Grasso P (1991). Chronic nitrosamine ingestion in 1040 rodents: The effect of the choice of nitrosamine, the species studied, and the age of starting exposure. *Cancer Research* 51: 6470-6491.

Hecht SS, Lijinsky W, Kovatch RM, Chung FL, Saavedra JE (1989). Comparative tumorigenicity of N-nitroso-2-hydroxymorpholine, N-nitrosodiethanolamine and N-nitrosomorpholine in A/J mice and F344 rats. *Carcinogenesis* 10: 1475 – 1477.

Habs M, Schmähl (1980). Synergistic effect of N-nitroso compounds in experimental long-term carcinogenesis studies. *Oncology* 37: 259-265.

Hoos A, Habst M, Schmähl D (1985). Comparison of liver tumor frequencies after intermittent oral administration of different doses of N-nitrosopyrrolidine in Sprague-Dawley rats. *Cancer Letters* 26: 77-82.

Konishi Y, Kondo H, Ikeda T, Kawabata A, Shoji Y, Deda A (1978). Effect of dose on the carcinogenic activity of orally administered N-bis(2-hydroxypropyl)nitrosamine in rats. *GANN* 69: 573-577.

Lijinsky W, Taylor HW (1976). The effect of substituents on the carcinogenicity of N-nitrosopyrrolidine in Sprague-Dawley Rats. *Cancer Research* 36: 1988-1990.

Lijinsky W, Reuber MD (1981). Carcinogenic effect of nitrosopyrrolidine, nitrosopiperidine and nitrosohexamethyleneimine in Fischer rats. *Cancer Letters* 12: 99-103.

Lijinsky W, Reuber MD, Riggs CW (1981). Dose response studies of carcinogenesis in rats by nitrosodiethylamine. *Cancer Research* 41: 4997-5003 (1981).

Lijinsky W, Reuber MD (1982). Comparative carcinogenesis by nitrosomorpholines, nitrosooxazolidines and nitrosotetrahydrooxazine in rats. *Carcinogenesis* 3: 911-915.

Lijinsky W, Reuber MD, Riggs CW (1983). Carcinogenesis by combinations of nitroso compounds in rats. *Food Chem. Toxicol* 21: 601 – 605.

Lijinsky W, Reuber MD (1984). Carcinogenesis in rats by nitrosodimethylamine and other nitrosomethylalkylamines at low doses. *Cancer Letters* 22: 83-88

Lijinsky W, Saavedra JE, Reuber MD (1984). Carcinogenesis in rats by some hydroxylated acyclic nitrosamines. *Carcinogenesis* 5: 167-170.

Lijinsky W, Kovatch RM (1985). Induction of liver tumors in rats by nitrosodiethanolamine at low doses. *Carcinogenesis* 6: 1679 – 1681.

Lijinsky W, Kovatch M, Riggs CV (1987) Carcinogenesis by nitrosodialkylamines and azoxyalkanes given by gavage to rats and hamsters. *Cancer Res.* 47: 3968-3972.

Lijinsky W, Kovatch RM, Riggs CW, Walters PT (1988). Dose response study with N-nitrosomorpholine in drinking water of F-344 rats. *Cancer Research* 48: 2089-2095.

Michejda CJ, Kroeger-Koepke MB, Kovatch RM (1986). Carcinogenic effects of sequential administration of two nitrosamines in Fischer 344 rats. *Cancer Research* 46: 2252-2256.

Peto R, Gray R, Brantom P, Grasso P (1991). Effects on 4080 rats of chronic ingestion of N-Nitrosodiethylamine or N-Nitrosodimethylamine: A Detailed Dose-Response Study. *Cancer Res* 51: 6415-6451.

Preussmann R, Habs M, Habs H, Schmähl D (1982). Carcinogenicity of N-nitrosodiethanolamine in rats at five different dose levels. *Cancer Res* 42: 5167 – 5171.

Preussmann R, Schmähl D, Eisenbrand G (1977). Carcinogenicity of N-Nitrosopyrrolidine: Dose-Response Study in Rats. *Z. Krebsforsch.* 90: 161-166.

Terao K, Alkawa T, Kera K (1978). A synergistic effect of nitrosodimethylamine on sterigmatocystin carcinogenesis in rats. *Fd Cosmet Toxicol* 16: 591-596.

Zerban H, Preussmann R, Bannasch P (1988). Dose-time relationship of the development of preneoplastic liver lesions induced in rats with low doses of N-nitrosodiethanolamine. *Carcinogenesis* 9: 607 - 610.

Annex II

Dose descriptors derived from a BMD analysis and ranking information:

The SCCS applied the Benchmark Dose (BMD) Approach to derive dose descriptors for the eight NOCs identified above from available dose response studies, where applicable. The BMD approach has been increasingly used and recommended (EFSA, 2009, 2011) to calculate a point-of departure for risk assessment. It represents a dose level, derived from the estimated dose-response curve, associated with a specific change in the response defined through a Benchmark Response level, denoted BMR. For cancer incidence, it has been set by default equal to 10% extra risk over background. Ideally the BMD uses all available dose-response data of a study and fits a set of mathematical models. The lower one-sided confidence bound BMDL accounts for the statistical uncertainty in the data (with the statistical certainty level of 95%) and is used as a point of departure.

The BMD approach requires however a sufficient quality of the dose response data in terms of the number of dose groups. Therefore, it cannot be applied to studies with only two groups (a control and one dose group). It may also not be applied to studies with only three groups (a control and two dose groups) since in that case not all models available for the BMD approach can be fitted such that a statistically sound BMDL can be calculated. Therefore the BMD analysis of the eight NOCs under consideration was restricted to those data sets where this criterion applied.

This Annex reports details of the BMD Analysis and those dose descriptors which represent the BMD approach in Table 8.

Table NDMA.1: Results of dose-response modelling of the NDMA data from Peto et al. (1991) on Colworth/Wistar male rats exposed by drinking water with fatal liver neoplasm incidence as endpoint (liver cells, bile ducts, mesenchyme, Kupffer cells, unknown) for the low dose range (control and 6 lowest dose groups), defined by Peto et al. (1991) as of lower than 1.1 ppm, for all models available in US EPA Software BMDS 2.1.1 and larger dose ranges for selected models (see lines with No. of Parameters equal to 9 or 16).

Model	Number of Parameters/ -loglikelihood ^{a)}		Model Acceptance p-value	BMD10 ppm	BMDL10 ppm
Full	7	76.4			
Reduced	1	84.3			
Probit	2	80.6	0.15	1.2	0.84
	16		<10 ⁻⁸	1.50	1.37
	9				
LogProbit	3	76.9	0.91	3.5	0.64
	16		0.054	1.46	1.24
Logistic	2	80.6	0.14	1.2	0.90
	16		<10 ⁻⁶	1.59	1.46
LogLogistic	3	76.9	0.90	3.1	0.65
	16		0.074	1.44	1.22
	9		0.65	1.94	1.63
Weibull	3	76.9	0.90	3.0	0.65

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	16		0.004	1.00	0.79
Gamma	3	76.9	0.90	2.9	0.66
	16		0.01	1.22	0.96
MS CA	2	79.4	0.30	1.1	0.63
Slope factor =	0.16				
	16		no fit possible		
Quantal Linear	2	79.4	same as MS CA		
	16		<0.05	0.50	0.45

bw: body weight; BMD: benchmark dose; BMDL10: 95 % lower confidence limit of the benchmark dose with 10 % extra risk, MS CA: Multistage Cancer model

^{a)} the negative loglikelihood value is reported only for the data set of the control group and the 6 lowest dose groups used for ranking

The BMDL value of 0.64 ppm was chosen as the BMDL10 of this study. Based on the conversion of the dosing in ppm to mg/kg bw/d calculated by OBrien et al. (2006) this value was converted by the factor of 24 into the **BMDL10 of 0.027 mg/kg bw/d**.

Table NDMA.2: Results of dose-response modelling of the NDMA data from Peto et al. (1991) on Colworth/Wistar female rats exposed by drinking water with fatal liver neoplasm incidence as endpoint (liver cells, bile ducts, mesenchyme, Kupffer cells, unknown) for the low dose range (control and 6 lowest dose groups) defined by Peto et al. (1991) as of lower than 1.1 ppm, for all models available in US EPA Software BMDS 2.1.1 and larger dose ranges for selected models (see lines with No. of Parameters equal to 9 or 16).

Model	Number of Parameters -loglikelihood ^{a)}		Model Acceptance p-value	BMD10 ppm	BMDL10 ppm
Full	7	66.7			
Reduced	1	77.3			
Probit	2	71.0	0.14	1.03	0.79
	16				
	9				
LogProbit	3	68.1	0.59	1.13	0.56
	16		0.004	0.57	0.45
Logistic	2	77.1	0.11	1.05	0.83
	16				
LogLogistic	3	68.1	0.56	1.10	0.58
	16		0.03	0.60	0.48
	9				
Weibull	3	68.2	0.56	1.10	0.59
	16		<10 ⁻⁵	0.35	0.27
Gamma	ND				

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MS CA	2	68.6	0.57	0.92	0.59
Slope factor =	0.17				
	16		<10 ⁻⁵	0.31	0.28
Slope factor =	0.36				
Quantal Linear	2	68.6	same as MS CA		

bw: body weight; BMD: benchmark dose; BMD₁₀: 95 % lower confidence limit of the benchmark dose with 10 % extra risk; MS CA: Multistage Cancer model,

ND: model could not be fitted

^{a)} the negative loglikelihood value is reported only for the data set of the control group and the 6 lowest dose groups used for ranking

The BMDL value of 0.56 ppm was chosen as the BMDL₁₀ of this study. Based on the conversion of the dosing in ppm to mg/kg bw/d calculated by OBrien et al. (2006) this value was converted by the factor of 13.8 into the **BMDL₁₀ of 0.041 mg/kg bw/d.**

Table NDEA.1: Results of dose-response modelling of the NDEA data from Peto et al. (1991) on Colworth/Wistar male rats exposed by drinking water with fatal liver neoplasm incidence as endpoint (liver cells, bile ducts, mesenchyme, Kupffer cells, unknown) for the low dose range (control and 6 lowest dose groups), defined by Peto et al. (1991) as of lower than 1.1 ppm, for all models available in US EPA Software BMDS 2.1.1 and larger dose ranges for selected models (see lines with No. of Parameters equal to 9 or 16).

Model	Number of Parameters	-loglikelihood ^{a)}	Model Acceptance p-value	BMD ₁₀ ppm	BMDL ₁₀ ppm
Full	7	75.8			
Reduced	1	91.0			
Probit	2	81.6	0.045	0.89	0.71
	16				
	9				
LogProbit	3	79.1	0.16	0.79	0.42
	16		0.07	0.74	0.50
Logistic	2	81.9	0.03	0.92	0.75
	16				
LogLogistic	3	79.1	0.16	0.79	0.44
	16		0.13	0.74	0.52
	9			0.84	0.50
Weibull	3	79.1	0.16	0.79	0.45
	16		0.16	0.69	0.48
Gamma	ND				
MSCA	2	79.7	0.16	0.72	0.48
Slope factor =	0.21				
	16		0.03	1.11	0.98
Slope factor = 0.10					
Quantal Linear 2	same as MS CA				

bw: body weight; BMD: benchmark dose; BMDL₁₀: 95% lower confidence limit of the benchmark dose with 10 % extra risk; MS CA: Multistage Cancer model

ND: model could not be fitted

^{a)} the negative loglikelihood value is reported only for the data set of the control group and the 6 lowest dose groups used for ranking

The BMDL value of 0.42 ppm was chosen as the BMDL₁₀ of this study. Based on the conversion of the dosing in ppm to mg/kg bw/d calculated by O'Brien et al. (2006) this value was converted by the factor of 24 into the **BMDL₁₀ of 0.018 mg/kg b.w./d.**

Table NDEA.2: Results of dose-response modelling of the NDEA data from Peto et al. (1991) on Colworth/Wistar female rats exposed by drinking water with fatal liver neoplasm incidence as endpoint (liver cells, bile ducts, mesenchyme, Kupffer cells, unknown) for low dose range (control and 6 lowest dose groups) defined by Peto et al. (1991) as of lower than 1.1 ppm, for all models available in US EPA Software BMDS 2.1.1 and larger dose ranges for selected models (see lines with No. of Parameters equal to 9 or 16).

Model	Number of Parameters	-loglikelihood ^{a)}	Model Acceptance p-value	BMD10 ppm	BMDL10 ppm
Full	7	69.4			
Reduced	1	116.1			
Probit	2	69.4	0.87	0.64	0.57
	16				
	9				
LogProbit	3	70.0	0.57	0.63	0.49
	16		<10 ⁻⁵	0.33	0.25
Logistic	2	69.4	0.88	0.70	0.62
	16				
LogLogistic	3	69.7	0.66	0.64	0.49
	16		<10 ⁻⁵	0.33	0.25
	9			0.56	0.11
Weibull	3	69.6	0.69	0.64	0.49
	16		<10 ⁻¹¹	0.18	0.13
Gamma	ND				
MultiStage Ca	2	70.6	0.52	0.54	0.46
Slope factor = 0.22					
	16		<10 ⁻¹⁵	0.37	0.41
Slope factor = 0.29					
Quantal Linear	2	77.7	<0.05	0.42	0.31

bw: body weight; BMD: benchmark dose; BMDL₁₀: 95% lower confidence limit of the benchmark dose with 10 % extra risk, MS CA: Multistage Cancer model

ND: model could not be fitted

a) the negative loglikelihood value is reported only for the data set of the control group and the 6 lowest dose groups used for ranking

The BMDL value of 0.49 ppm was chosen as the BMDL₁₀ of this study. Based on the conversion of the dosing in ppm to mg/kg bw/d calculated by OBrien et al. (2006) this value was converted by the factor of 13.8 into the **BMDL₁₀ of 0.036 mg/kg b.w./d.**

Table NDEA.3: Results of dose-response modelling of the NDEA data from Berger et al. (1987, 1990) on SD male rats exposed by drinking water with liver tumour incidence as endpoint for all models available in US EPA Software BMDS 2.1.1. for all 4 groups (Incidence: control: 3/500, 0.01:2/80;0.032:2/80;0.1:36/80)

Model	Number of Parameters	-loglikelihood	Model Acceptance p-value	BMD ₁₀ mg/kg bw/d	BMDL ₁₀ mg/kg bw/d
Full	4	95.5			
Reduced	1	166.9			
Probit	2	96.0	0.66	0.052	0.046
LogProbit	3	96.6	0.15	0.047	0.034
Logistic	2	96.1	0.59	0.058	0.052
LogLogistic	3	96.4	0.19	0.048	0.034
Weibull	3	96.3	0.21	0.048	0.034
Gamma	ND				
MS CA	2	96.5	0.17	0.043	0.033
Slope factor =	3.03				
Quantal Linear	2	101.9	0.002	0.025	0.019

bw: body weight; BMD: benchmark dose; BMDL₁₀: 95% lower confidence limit of the benchmark dose with 10 % extra risk, MS CA: Multistage Cancer model

Table NDELA.1: Results of dose-response modelling of the NDELA data from Berger et al. (1987, 1990) on SD male rats exposed by drinking water with liver tumour incidence as endpoint for all models available in US EPA Software BMDS 2.1.1 for all 4 groups (Incidence: control: 3/500, 0.2:2/80;0.63:1/80;2.0:6/80)

Model	Number of Parameters	-loglikelihood	Model Acceptance p-value	BMD ₁₀ mg/kg bw/d	BMDL ₁₀ mg/kg bw/d
Full	4	61.4			

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Reduced	1	54.4			
Probit	2	55.2	0.45	2.40	1.81
LogProbit	3	55.2	0.19	4.11	1.74
Logistic	2	55.2	0.44	2.33	1.83
LogLogistic	3	55.2	0.22	3.44	1.74
Weibull	3	55.2	0.22	3.37	1.74
Gamma	3	55.2	0.22	3.39	1.74
MS CA	2	55.2	0.46	3.20	1.76
Slope factor =	0.33				

Quantal Linear = MS CA

bw: body weight; BMD: benchmark dose; BMDL₁₀: 95% lower confidence limit of the benchmark dose with 10 % extra risk, MS CA: Multistage Cancer model

Table NDELA.2: Results of dose-response modelling of the NDELA data from Preussmann et al. (1982) on SD male rats exposed by drinking water with liver tumour incidence as endpoint for all models available in US EPA Software BMDS 2.1.1 for 4 groups (Incidence: control: 0/88, 1.5:7/72;6:43/72;25:33/36) Without the two largest dose groups

Model	Number of Parameters	-loglikelihood	Model Acceptance p-value	BMD10 mg/kg bw/d	BMDL10 mg/kg bw/d
Full	4	81.8			
Reduced	1	165.9			
Probit	2	105.6	<10 ⁻¹⁰	2.81	2.34
LogProbit	3	82.3	0.64	1.41	0.98
Logistic	2	103.0	<10 ⁻⁹	2.45	2.04
LogLogistic	3	82.1	0.74	1.41	0.95
Weibull	3	84.5	0.07	0.95	0.55
Gamma	ND				
MS CA	2	84.5	0.14	0.89	0.73
Slope factor =	0.14				

Quantal Linear = MS CA

bw: body weight; BMD: benchmark dose; BMDL₁₀: 95% lower confidence limit of the benchmark dose with 10 % extra risk, MS CA: Multistage Cancer model

Table NPYR.1: Results of dose-response modelling of the NPYR data from Berger et al. (1987, 1990) on SD male rats exposed by drinking water with liver tumour incidence as endpoint for all models available in US EPA Software BMDS 2.1.1 for all 4 groups (Incidence: control: 3/500, 0.04:1/80;0.133:4/80;0.4:17/80)

Model	Number of Parameters -loglikelihood		Model Acceptance p-value	BMD10 mg/kg bw/d	BMDL10 mg/kg bw/d
Full	4	81.0			
Reduced	1	109.3			
Probit	2	81.5	0.56	0.29	0.25
LogProbit	3	81.0	0.84	0.22	0.15
Logistic	2	81.9	0.41	0.30	0.27
LogLogistic	3	81.0	0.98	0.23	0.16
Weibull	3	81.0	0.93	0.24	0.17
Gamma	ND				
MS CA Slope factor =	2	81.0 0.60	0.83	0.25	0.17
Quantal Linear 2	81.7		0.46	0.21	0.15

bw: body weight; BMD: benchmark dose; BMDL₁₀: 95% lower confidence limit of the benchmark dose with 10 % extra risk, MS CA: Multistage Cancer model

For N-nitrosodimethylamine **NDMA** data sets from a total of 5 studies were available. Only two studies fulfilled the quality criteria required for applying the BMD approach. One of the studies (Arai et al., 1979) reported the incidence of hepatocellular carcinoma separately for males and female Wistar rats in a dose range up to 10 ppm. A BMDL₁₀ was calculated to 2.0 ppm in standard diet for hepatocellular carcinoma in males and 0.5 ppm in females, respectively. The study with highest quality available for NDMA was that of Peto et al. (1991) where ample dose-response data were available both for male and female Wistar rats. Restricting the BMD analysis on the lower dose range (up to 1.056 ppm) the BMDL for males and females was determined as 0.64 and 0.56 ppm, respectively. Those values were transformed using the calculations of O'Brien et al. (2006) into 0.027 and 0.041 mg/kg b.w./d, respectively. Therefore, a BMDL₁₀ = 0.027 mg/kg b.w./d was established as dose descriptor for NDMA.

For N-nitrosodiethylamine (**NDEA**) data sets from total of 6 studies were available. Only two studies (Peto et al. (1991), Berger et al. (1987, 1990)) fulfilled the quality criteria required for applying the BMD approach. When restricting the study of Peto et al. (1991) where ample dose-response data were available both for male and female Wistar rats on the lower dose range (up to 1.056 ppm) the BMDL₁₀ for males and females was determined as 0.42 ppm both for males and females. Those values were transformed using the calculations of O'Brien et al. (2006) into 0.018 and 0.030 mg/kg bw/d respectively. The

study of Berger et al. (1987, 1990) gave a BMDL = 0.034 mg/ kg bw/d as dose descriptor for NDEA. The lowest BMDL10 of 0.018 mg/kg bw/d was established as dose descriptor for NDEA.

Only sparse data were available for N-nitrosomorpholine (**NMOR**) (4 studies), for N-nitroso-N-methyl-N-dodecylamine (1 study) and for N-nitroso-bis(2-hydroxypropyl)amine (**NBHPA**, 2 studies). None of those studies allowed to application of the BMD approach. Using the data of the study of Lijinsky et al (1988) for NMOR the SCCS derived a BMDL-10 value of 1.7 mg/L (the minimum BMDL-10 of all acceptable models fitted) and using the conversion to body weight indicated in the publication a BMDL-10 of 0.70 mg/kg b.w/d was calculated for hepatocellular carcinoma incidence.

For N-nitrosopyrrolidine (**NPYR**) data from a total of 8 studies were available. However only the study Berger et al. (1987, 1990) was amenable to a BMD analysis which resulted in a BMDL10 = 0.16 mg/kg/d.

For N-nitrosodiethanolamine (**NDELA**) data sets from six studies of different quality were available. There were three studies for which the BMD approach could be applied: Preussman et al. (1982), Berger et al. (1987, 1990) and Zerban et al. (1988). The study of Zerban et al. (1988) investigated a control and five dose groups (0.2, 0.63, 1.5, 6, and 25 mg/kg bw/d) for hepatocellular carcinomas and hepatic nodules and their incidence over a time period of 24 months with interim sacrifice every 3 months after 12 months such that only a small portion of animals was available at end of study for calculating a total incidence. Therefore, the BMDL10 = 1.6 mg/kg/d derived from that study was not considered as reliable as the values obtained from the other two studies which had much higher numbers of evaluable animals. Not all dose-response data of the Preussmann et al. (1982) study could be used for the BMD analysis since the incidence of liver tumours reached at the dose of 25 mg/kg bw/d almost the level of 100%. Then restricting the BMD analysis to doses between 1.5 and 25 mg/kg bw/d a BMDL10=0.55 mg/kg bw/d was calculated. The analysis of data of the other high quality study of Berger et al. (1987, 1990) where the doses range as low as 0.2 mg/kg b.w./day yielded BMDL10= 1.74 mg/kg b.w./day, larger by a factor of three. Tentatively combining the data of both studies which were based on similar designs a joint BMDL10 of 0.73 mg/kg/d was calculated which supported the data of Zerban et al. (1988) used for the ranking.